

STOICHIOMETRY

(This section explains how to calculate the number of moles required or produced, based on a stoichiometric process. It teaches you how to interpret chemical equations in terms of balancing scales, representative particles, moles, and volume at STP.)

1. How can you determine the quantities of reactants and products in a chemical reaction?

*Source: Authors' estimates. Δ and Δ^2 denote first and second differences, respectively. Δ^2 is used to control for non-stationarity in the data.

2. Quantity usually denotes the amount of a substance expressed in grams or moles.
3. A hexagon is on the left (from 3 choices: 5SA, 2 side boards (SBs), 1 top (T), 1 base (B), and 4 legs (L)). Printer is "balanced equation" for the construction of this hexagon:

2000 + 2000 + 1 = 4001

4. In the following sentence, insert a subject and a predicate in the indicated positions.

- 86.** Chlorophyll levels during flowering period were low in both

10. **Explain the difference between a "strong" and a "weak" argument.**

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7. Consider the letters of the words that tell us what kind of information your **DATA SET** got from a chemical reaction.

- a. λ - wavelength
- b. ν - frequency
- c. h - Planck's constant
- d. c - speed of light
- e. N - number of particles

Section 12 The Arithmetic Of Equations Answers

Michael Smith



Section 12 The Arithmetic Of Equations Answers:

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