

# SECTION 1

## Reinforcement

## Chemical Changes

Key

Directions: Use the equations below to answer the following questions.



1. What are the reactants in this chemical reaction?

Zinc + Sulfur

2. What is the product?

Zinc Sulfide

3. What is the physical state of both the reactants and the products?

Solid

4. According to the law of conservation of mass, if the total mass of the product in this chemical reaction is 14 g, what must the combined masses of the reactants be?

14 g



5. What name describes the product in this reaction?

Water

6. What names describe the reactants?

Hydrogen + Oxygen

7. What are the physical states of the reactants in this reaction?

gas

8. What is the physical state of the product?

liquid

9. What do the coefficients tell you about the ratio of the reactants?

It takes twice as many Hydrogen as it does Oxygen

10. How many units of the product are produced?

2

Directions: Write chemical equations for the following reactions.

11. One unit of methane gas,  $\text{CH}_4$ , plus two units of oxygen gas,  $\text{O}_2$ , produce one unit of carbon dioxide gas,  $\text{CO}_2$ , and two units of liquid water.



12. One unit of aqueous aluminum sulfate,  $\text{Al}_2(\text{SO}_4)_3$ , plus three units of aqueous barium chloride,  $\text{BaCl}_2$ , yield two units of aqueous aluminum chloride,  $\text{AlCl}_3$ , plus three units of solid barium sulfate,  $\text{BaSO}_4$ .



13. Two units of solid sodium,  $\text{Na}$ , plus one unit of chlorine gas produce two units of sodium chloride,  $\text{Cl}_2$ , a solid.



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