

Limiting Reagent & Percent Yield Practice Worksheet

1. When copper (II) chloride reacts with sodium nitrate, copper (II) nitrate and sodium chloride are formed.

a. Write the balanced equation for the reaction given above:



- b. If 15 grams of copper (II) chloride react with 20 grams of sodium nitrate, how much sodium chloride can be formed?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.362 \text{ g CuCl}_2} \times \frac{2 \text{ mol NaCl}}{1 \text{ mol CuCl}_2} \times \frac{58.443 \text{ g NaCl}}{1 \text{ mol NaCl}} = 13 \text{ g NaCl}$$

$$20 \text{ g NaNO}_3 \times \frac{1 \text{ mol NaNO}_3}{84.997 \text{ g NaNO}_3} \times \frac{2 \text{ mol NaCl}}{2 \text{ mol NaNO}_3} \times \frac{58.443 \text{ g NaCl}}{1 \text{ mol NaCl}} = 13.8 \text{ g NaCl}$$

Only 13 g can be formed

- c. What is the limiting reagent for the reaction in #2? **copper (II) chloride**

- d. How many grams of copper (II) nitrate is formed?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.452 \text{ g CuCl}_2} \times \frac{1 \text{ mol Cu(NO}_3)_2}{1 \text{ mol CuCl}_2} \times \frac{187.467 \text{ g Cu(NO}_3)_2}{1 \text{ mol Cu(NO}_3)_2} = 20.9 \text{ g Cu(NO}_3)_2$$

- e. How much of the excess reagent is left over in this reaction?

$$15 \text{ g CuCl}_2 \times \frac{1 \text{ mol CuCl}_2}{134.452 \text{ g CuCl}_2} \times \frac{2 \text{ mol NaNO}_3}{1 \text{ mol CuCl}_2} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 19 \text{ g NaNO}_3$$

$$20 \text{ g NaNO}_3 - 19 \text{ g NaNO}_3 = 1 \text{ g NaNO}_3$$

- f. If 11.3 grams of sodium chloride are formed in the reaction described in problem #2, what is the percent yield of this reaction?

$$\text{Percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100 \text{ percent}$$

$$11.3/13.0 \times 100\% = 86.9\%$$

2. When lead (II) nitrate reacts with sodium iodide, sodium nitrate and lead (II) iodide are formed.

a. Balance the following equation:



- b. If I start with 25.0 grams of lead (II) nitrate and 15.0 grams of sodium iodide, how many grams of sodium nitrate can be formed?

$$25.0 \text{ g Pb(NO}_3)_2 \times \frac{1 \text{ mol Pb(NO}_3)_2}{331.214 \text{ g Pb(NO}_3)_2} \times \frac{2 \text{ mol NaNO}_3}{1 \text{ mol Pb(NO}_3)_2} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 12.8 \text{ g NaNO}_3$$

$$15.0 \text{ g NaI} \times \frac{1 \text{ mol NaI}}{149.89 \text{ g NaI}} \times \frac{2 \text{ mol NaNO}_3}{2 \text{ mol NaI}} \times \frac{84.997 \text{ g NaNO}_3}{1 \text{ mol NaNO}_3} = 8.51 \text{ g NaNO}_3$$

8.51 g NaNO₃ can be made

Section 12 3 Limiting Reagent And Percent Yield

Answers

Kaplan Test Prep



Section 12 3 Limiting Reagent And Percent Yield Answers:

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Table of Contents Section 12 3 Limiting Reagent And Percent Yield Answers

1. Understanding the eBook Section 12 3 Limiting Reagent And Percent Yield Answers
 - The Rise of Digital Reading Section 12 3 Limiting Reagent And Percent Yield Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Section 12 3 Limiting Reagent And Percent Yield Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section 12 3 Limiting Reagent And Percent Yield Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section 12 3 Limiting Reagent And Percent Yield Answers
 - Personalized Recommendations
 - Section 12 3 Limiting Reagent And Percent Yield Answers User Reviews and Ratings

- Section 12 3 Limiting Reagent And Percent Yield Answers and Bestseller Lists
- 5. Accessing Section 12 3 Limiting Reagent And Percent Yield Answers Free and Paid eBooks
 - Section 12 3 Limiting Reagent And Percent Yield Answers Public Domain eBooks
 - Section 12 3 Limiting Reagent And Percent Yield Answers eBook Subscription Services
 - Section 12 3 Limiting Reagent And Percent Yield Answers Budget-Friendly Options
- 6. Navigating Section 12 3 Limiting Reagent And Percent Yield Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Section 12 3 Limiting Reagent And Percent Yield Answers Compatibility with Devices
 - Section 12 3 Limiting Reagent And Percent Yield Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section 12 3 Limiting Reagent And Percent Yield Answers
 - Highlighting and Note-Taking Section 12 3 Limiting Reagent And Percent Yield Answers
 - Interactive Elements Section 12 3 Limiting Reagent And Percent Yield Answers
- 8. Staying Engaged with Section 12 3 Limiting Reagent And Percent Yield Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section 12 3 Limiting Reagent And Percent Yield Answers
- 9. Balancing eBooks and Physical Books Section 12 3 Limiting Reagent And Percent Yield Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section 12 3 Limiting Reagent And Percent Yield Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Section 12 3 Limiting Reagent And Percent Yield Answers
 - Setting Reading Goals Section 12 3 Limiting Reagent And Percent Yield Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Section 12 3 Limiting Reagent And Percent Yield Answers
 - Fact-Checking eBook Content of Section 12 3 Limiting Reagent And Percent Yield Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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