

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH values and concentrations based on the pH and  $K_a$  or  $K_b$  and vice versa.
3. Calculate the pH of weak acids and bases.

## Vocabulary

- |                   |                             |                    |
|-------------------|-----------------------------|--------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base  |
| 4. autoionization | 5. acid rain                | 6. buffer solution |
| 7. autoionization | 8. acid rain                | 9. pH              |

## Key Equation

- |                                    |                                    |
|------------------------------------|------------------------------------|
| 1. $\text{pH} = -\log[\text{H}^+]$ | 2. $\text{pH} = -\log[\text{H}^+]$ |
| 3. $\text{pH} = -\log[\text{H}^+]$ | 4. $\text{pH} = -\log[\text{H}^+]$ |

## Part A Completion

Write the equation and/or the chemical formula(s) of the species involved in the reaction. Write the chemical formula(s) of the species involved in the reaction.

1. Write the equation and/or the chemical formula(s) of the species involved in the reaction.
2. Write the equation and/or the chemical formula(s) of the species involved in the reaction.
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17. Write the equation and/or the chemical formula(s) of the species involved in the reaction.
18. Write the equation and/or the chemical formula(s) of the species involved in the reaction.
19. Write the equation and/or the chemical formula(s) of the species involved in the reaction.
20. Write the equation and/or the chemical formula(s) of the species involved in the reaction.

1.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
2.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
3.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
4.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
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20.  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$

## Section 19 2 Hydrogen Ions And Acidity Answer Key

**RJ Alexander**



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