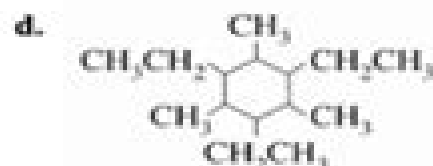


# CHAPTER 21 Hydrocarbons

1. Use the IUPAC rules to name the following alkanes.



2. Draw the structure of each of the following alkanes.

a. 4-propyloctane

b. 3,4-diethylhexane

c. 2,2,4,4-tetramethylhexane

d. 1-ethyl-3-methyl-2-propylcyclopentane

3. Calculate the number of hydrogen atoms in each of the following alkanes.

a. heptane

b. cyclooctane

4. Calculate the molecular mass of a 22-carbon branched-chain alkane.

5. Chemists can analyze the composition of hydrocarbons by reacting them with copper oxide. The reaction converts carbon into carbon dioxide and hydrogen into water. Suppose 29 g of a hydrocarbon reacts to produce 88 g of  $\text{CO}_2$  and 45 g of  $\text{H}_2\text{O}$ .

a. What are the masses of carbon and hydrogen in the hydrocarbon?

b. What is the empirical formula of the hydrocarbon?

c. If the hydrocarbon's molecular mass is 58 amu, what is its molecular formula?

6. Carbon has an electronegativity of 2.5. Hydrogen has an electronegativity of 2.2. Use these values to decide whether each of the following bonds is polar or nonpolar.

a. C-C

b. C-H

c. H-H

7. The combustion of a saturated hydrocarbon releases 657 kJ per mole of  $-\text{CH}_2-$  groups and 779 kJ per mole of  $-\text{CH}_3$  groups in the hydrocarbon. How much energy is released by the combustion of 1.00 L of liquid tetradecane (molecular formula  $\text{C}_{14}\text{H}_{30}$ ), a major component of kerosene? The density of tetradecane is 0.764 g/mL.

## Section 22 Hydrocarbons Answer Key

**Clifford Lane**



## Section 22 Hydrocarbons Answer Key:

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group II metals uses of halogens and their compounds The Halogenoalkanes Quiz Questions PDF e Book Chapter 14 interview questions and answers on Halogenoalkanes uses of halogenoalkanes elimination reactions nucleophilic substitution in halogenoalkanes and nucleophilic substitution reactions The Hydrocarbons Quiz Questions PDF e Book Chapter 15 interview questions and answers on Introduction to alkanes sources of alkanes addition reactions of alkenes alkane reaction alkenes and formulas The Introduction to Organic Chemistry Quiz Questions PDF e Book Chapter 16 interview questions and answers on Organic chemistry functional groups organic reactions naming organic compounds stereoisomerism structural isomerism and types of organic reactions The Ionic Equilibria Quiz Questions PDF e Book Chapter 17 interview questions and answers on Introduction to ionic equilibria buffer solutions equilibrium and solubility indicators and acid base titrations pH calculations and weak acids The Lattice Energy Quiz Questions PDF e Book Chapter 18 interview questions and answers on Introduction to lattice energy ion polarization lattice energy value atomization and electron affinity Born Haber cycle and enthalpy changes in solution The Moles and Equations Quiz Questions PDF e Book Chapter 19 interview questions and answers on Amount of substance atoms molecules mass chemical formula and equations gas volumes mole calculations relative atomic mass solutions and concentrations The Nitrogen and Sulfur Quiz Questions PDF e Book Chapter 20 interview questions and answers on Nitrogen gas nitrogen and its compounds nitrogen and gas properties ammonia ammonium compounds environmental problems caused by nitrogen compounds and nitrate fertilizers sulfur and oxides sulfuric acid and properties and uses of sulfuric acid The Organic and Nitrogen Compounds Quiz Questions PDF e Book Chapter 21 interview questions and answers on Amides in chemistry amines amino acids peptides and proteins The Periodicity Quiz Questions PDF e Book Chapter 22 interview questions and answers on Acidic oxides basic oxides aluminum oxide balancing equation period 3 chlorides balancing equations reactions with chlorine balancing equations reactions with oxygen bonding nature of period 3 oxides chemical properties of chlorine chemical properties of oxygen chemical properties periodicity chemistry periodic table chemistry oxides chlorides of period 3 elements electrical conductivity in period 3 oxides electronegativity of period 3 oxides ionic bonds molecular structures of period 3 oxides oxidation number of oxides oxidation numbers oxides and hydroxides of period 3 elements oxides of period 3 elements period III chlorides periodic table electronegativity physical properties periodicity reaction of sodium and magnesium with water and relative melting point of period 3 oxides The Polymerization Quiz Questions PDF e Book Chapter 23 interview questions and answers on Types of polymerization polyamides polyesters and polymer deductions The Rates of Reaction Quiz Questions PDF e Book Chapter 24 interview questions and answers on Catalysis collision theory effect of concentration reaction kinetics and temperature effect on reaction rate The Reaction Kinetics Quiz Questions PDF e Book Chapter 25 interview questions and answers on Reaction kinetics catalysts kinetics and reaction mechanism order of reaction rate constant  $k$  and rate of reaction The Redox Reactions and Electrolysis Quiz Questions PDF e Book Chapter 26 interview questions and answers on Redox reaction electrolysis

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## **Table of Contents Section 22 Hydrocarbons Answer Key**

1. Understanding the eBook Section 22 Hydrocarbons Answer Key
  - The Rise of Digital Reading Section 22 Hydrocarbons Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Section 22 Hydrocarbons Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Section 22 Hydrocarbons Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Section 22 Hydrocarbons Answer Key
  - Personalized Recommendations
  - Section 22 Hydrocarbons Answer Key User Reviews and Ratings
  - Section 22 Hydrocarbons Answer Key and Bestseller Lists
5. Accessing Section 22 Hydrocarbons Answer Key Free and Paid eBooks
  - Section 22 Hydrocarbons Answer Key Public Domain eBooks
  - Section 22 Hydrocarbons Answer Key eBook Subscription Services
  - Section 22 Hydrocarbons Answer Key Budget-Friendly Options
6. Navigating Section 22 Hydrocarbons Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Section 22 Hydrocarbons Answer Key Compatibility with Devices
  - Section 22 Hydrocarbons Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Section 22 Hydrocarbons Answer Key
  - Highlighting and Note-Taking Section 22 Hydrocarbons Answer Key
  - Interactive Elements Section 22 Hydrocarbons Answer Key
8. Staying Engaged with Section 22 Hydrocarbons Answer Key

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Section 22 Hydrocarbons Answer Key
- 9. Balancing eBooks and Physical Books Section 22 Hydrocarbons Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Section 22 Hydrocarbons Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Section 22 Hydrocarbons Answer Key
  - Setting Reading Goals Section 22 Hydrocarbons Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Section 22 Hydrocarbons Answer Key
  - Fact-Checking eBook Content of Section 22 Hydrocarbons Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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