

CHAPTER 6 REVIEW

The Periodic Law**SECTION 3****SHORT ANSWER** Answer the following questions in the space provided.

1. _____ When an electron is added to a neutral atom, energy is
 - (a) always absorbed.
 - (b) always released.
 - (c) either absorbed or released.
 - (d) neither absorbed nor released.
2. _____ The energy required to remove an electron from a neutral atom is the atom's
 - (a) electron affinity.
 - (b) electron energy.
 - (c) electronegativity.
 - (d) neither absorbed nor released.
3. From left to right across a period on the periodic table,
 - _____ a. electron affinity values tend to become more (negative or positive).
 - _____ b. ionization energy values tend to (increase or decrease).
 - _____ c. atomic radii tend to become (larger or smaller).
4. _____ a. Name the halogen with the least-negative electron affinity.
_____ b. Name the alkali metal with the highest ionization energy.
_____ c. Name the element in Period 3 with the smallest atomic radius.
_____ d. Name the Group 14 element with the largest electronegativity.
5. Write the electron configuration of the following:
 - a. Na

 - b. Na⁺

 - c. O

 - d. O²⁻

 - e. Co²⁺

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Liquids, Solutions, and Interfaces W. Ronald Fawcett, 2004-07 Fawcett chemistry University of California Davis introduces modern topics in solution chemistry to senior undergraduates and graduate students who have completed two semesters or three quarters of chemical thermodynamics and statistical mechanics

Journal of Sedimentary Petrology, 1992 **Academy and Literature** Charles Edward Cutts Birch Appleton, Charles Edward Doble, James Sutherland Cotton, Charles Lewis Hind, William Teignmouth Shore, Alfred Bruce Douglas, Ellis Ashmead-Bartlett, Thomas William Hodgson Crosland, 1892

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