#### Solving Thermodynamics Problems

Solving thermodynamic problems can be made significantly easier by using the following procedure:

- Summarize given data in own words, leave out unneeded information
- Clearly understand/identify what is being asked for draw a sketch showing interactions/states and identify a solution strategy. Keep in mind that a multistep approach may be the easiest
- 3. Define system boundaries, noting if it is an open or closed system
- 4. Fix the known states on a P-v/T-v/T-s diagram, using given information
- Apply conservation of mass to process
  - For a control mass/closed system:

$$m_{\nu} - m_{\nu} = 0$$

For a control volume/open system:

$$\frac{dm}{dt} = \sum \hat{m}_i - \sum \hat{m}_e$$

- 6. Apply conservation of energy to process (1st law of thermodynamics)
  - For a control mass/closed system:

$$\Delta E = Q - W$$
 (equilibrium process)

$$\frac{dE}{dt} = \dot{Q} - \dot{W}$$
 (rate form)

For a control volume/open system:

$$\frac{dE_{CV}}{dt} = Q^* - W^* + \sum m_i^* \left( h + \frac{V^2}{2} + gz \right)_{cc} - \sum m_o^* \left( h + \frac{V^2}{2} + gz \right)_{cc}$$

- Solve algebraically for desired quantity using combination of mass balance, energy balance, and definitions (like mass flow rate, volume, etc.)
- Perform heat transfer analysis to get Q, if necessary (use catalog of heat transfer or steady flow device in appendix)
- Perform work analysis to get W, if necessary (use catalog of work or steady flow device in appendix)
- 10. Find properties using Tables or models, such as ideal gas relationship
- 11. Substitute numbers into equation and solve for desired quantity

# **Solved Thermodynamic Problemscallenchapter15**

The Editors of REA, Ralph Pike

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