

# 8 • Electron Configurations & Periodicity

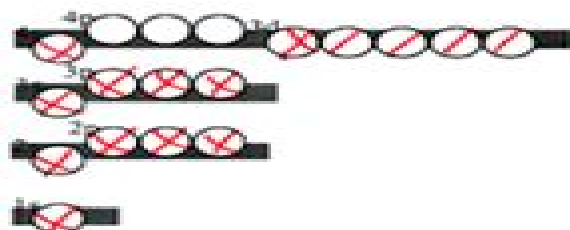
## WRITING ELECTRON CONFIGURATIONS

For each given element, fill in the orbital diagram and then write the electron configuration for the element.

1.	2.	3.	4.	5.	6.
Element: Ar # of e <sup>-</sup> s: <u>18</u>	Element: Mg # of e <sup>-</sup> s: <u>12</u>	Element: N # of e <sup>-</sup> s: <u>7</u>	Element: Li # of e <sup>-</sup> s: <u>3</u>	Element: P # of e <sup>-</sup> s: <u>15</u>	Element: Cl # of e <sup>-</sup> s: <u>17</u>

Write the electron configurations of each of these in **long form** and **short form**:

- Ar  $1s^2 2s^2 2p^6 3s^2 3p^4$   
Ar  $[Ne] 3s^2 3p^4$
- Mg  $1s^2 2s^2 2p^6 3s^2$   
Mg  $[Ne] 3s^2$
- N  $1s^2 2s^2 2p^3$   
N  $[He] 2s^2 2p^3$
- Li  $1s^2 2s^1$   
Li  $[He] 2s^1$
- P  $1s^2 2s^2 2p^6 3s^2 3p^3$   
P  $[Ne] 3s^2 3p^3$
- Cl  $1s^2 2s^2 2p^6 3s^2 3p^5$   
Cl  $[Ne] 3s^2 3p^5$



- Fill in the orbital diagram for the element, Fe, and write the electron configuration of Fe in the long and short form.



# Unit 6 Electron Configuration And Periodicity Review

## Answer Key

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you 23  $\text{Fe}^{2+} + \text{Co} \rightarrow \text{Co}^{2+} + \text{Fe}$   
oxidation  $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^{-}$  reduction  $\text{Fe}^{2+} + 2\text{e}^{-} \rightarrow \text{Fe}$  24 3  $\text{Ag} + \text{Ni} \rightarrow \text{Ni}^{3+} + 3\text{Ag}$  oxidation reduction 25  $\text{Cu}^{2+} + \text{Pb} \rightarrow \text{Pb}^{2+} + \text{Cu}$  oxidation reduction 26  $\text{O}_2$

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