

# Chemistry: Atom Unit Study Guide

24. Fill in the boxes below for the following concepts:

1. Isotope
2. Atomic Number
3. Mass Number
4. Proton
5. Neutron
6. Electron

25. Draw a Bohr model of the element

26. Complete the following table:

27. Draw a Bohr model of the element

28. Using the periodic table, determine the number of protons, neutrons, and electrons of  $^{13}\text{C}$ .

Name: \_\_\_\_\_ Period: \_\_\_\_\_

## Atom Unit Study Guide

**Matching:** Draw a line to match the words listed to the definitions.

- |                  |  |
|------------------|--|
| 1. Isotope       | — Atoms of the same element with different numbers of neutrons.                  |
| 2. Atomic Number | — The subscript particle has a positive charge.                                  |
| 3. Mass Number   | — The subscript particle has a negative charge.                                  |
| 4. Proton        | — The subscript particle has no charge (neutral).                                |
| 5. Neutron       | — The number indicates the total number of protons and neutrons in an atom.      |
| 6. Electron      | — This number indicates the number of protons in an atom and defines an element. |

### Circle the Blank.

**Word Bank:** Isotope, Neutrons, Empty Mass, Neutrons, Space, Protons, Number  
(Words can be used more than once)

- Where is most of an atom's mass found? \_\_\_\_\_
- An atom is made \_\_\_\_\_
- Which subatomic particle is the atomic number equal to? \_\_\_\_\_
- The atomic number is equal to \_\_\_\_\_ & \_\_\_\_\_ the combined subscripts that indicate all Neutrons in \_\_\_\_\_
- Isotopes of the same atom have the same \_\_\_\_\_ number but different \_\_\_\_\_ numbers.

### Calculate.

- Determine the number of protons, neutrons, and electrons of  $^{13}\text{C}$ .
- How many neutrons does the element Silver (Ag) have?

# Unit Study Guide Chemistry Answers

**M Planty**



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