

# AP Chemistry Unit 4 Review

## Redox Reactions and Electrochemistry

Honors Lecture Notes: Unit 6

### Oxidation Numbers

- 1) For any atom in its elemental form (including diatomics) the oxidation number is zero
- 2) For any monoatomic ion, the oxidation number is equal to the charge of the ion
- 3) Oxygen in a compound is always -2, except in peroxides, where it is -1
- 4) Hydrogen is +1, when bonded to a nonmetal, and -1 when bonded to a metal
- 5) For covalent compounds that do not contain O or H, the most electronegative element matches the charge of the negative ion it would create (i.e., Cl = -1, S = -2, N = -3, etc)
- 6) The total oxidation number for a compound must be 0 (for a neutral compound) or equal the total compound charge

If the oxidation number for an element becomes more negative, that element gained electrons and was reduced. If the oxidation number for an element becomes more positive, that element lost electrons and was oxidized.

### Simple Redox Reactions

Four of the most common redox reactions are listed below.

- 1) Solid metal in a metal solution. If the SRP of the ion in solution exceeds the SRP of the corresponding metal ion, the ion in solution will be reduced.



- 2) If a halogen gas is bubbled through a halide solution and the electronegativity of the gas exceeds that of the corresponding halide, the gas will be reduced.



- 3) If a metal is added to water, and the SRP of that metal's ions are less than that of the water, the water will be reduced. Most common with alkali/alkaline earth metals.



- 4) If a metal is added to any strong acid except nitric, and the SRP of that metal's ions are less than that of the  $\text{H}^+$  ions, the acid will be reduced, dissolving the metal.



### Redox Titrations

It would be a good idea to be familiar with (or even memorize) the following reduction half-reactions, which are often used in redox titrations. When doing a redox titration, make sure that the charges on both half-reactions balance and that the correct stoichiometric ratios are used during calculations.

- 1) Nitric Acid. The nitrate ion has a higher SRP than  $\text{H}^+$  ions, and can oxidize most metals.



- 2) Permanganate. The permanganate ion is a deep purple in color, and reduces into the colorless  $\text{Mn}^{2+}$  ion.



- 3) Dichromate. Dichromate ions (orange) will reduce to  $\text{Cr}^{3+}$  (varying colors, based on other species in solution).



### Electrochem Measurements

Current measures the rate at which charge (carried by electrons) is transferring. It is measure in Amperes, aka Coulombs/sec.

Voltage is a measure of electric potential. It is measured in Volts, aka Joules/Coulomb. Current will always flow from areas of high voltage to low voltage.

Faraday's Constant describes the amount of charge carried by 1 mole of electrons, and is equal to 96,500 C

### Reaction Favorability

A favored reaction is one in which the total cell potential is positive. A positive voltage also corresponds with a negative value for  $\Delta G$  given:  $\Delta G = -nFE$ , where:

$\Delta G$  is Gibbs free energy, in Joules       $F$  is Faraday's Constant  
 $n$  is moles of  $\text{e}^-$  transferred       $E$  is the full cell potential

Correspondingly, any cell with a positive voltage will always favor the products, and have a value of  $K$  greater than one.

# Unit 1chemical Reactions Review Sheet Answers

**John McMurry,Susan McMurry**



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Cambridge Primary Science Stage 6 Teacher's Resource Book with CD-ROM Fiona Baxter, Liz Dille, 2014-05-22

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**similarities between funzalushaka and nsfas bursary**

### **Unit 1chemical Reactions Review Sheet Answers :**

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