

Unit 2 Worksheet: St. Michael's Concentration

Part 2: Describing Molar Concentrations

1. Sodium chloride was dissolved in water to produce a 1.5M solution. Suppose that this concentration reflects about the same concentration as observed near the coast.
 What weight of sodium chloride is needed for a volume of 1.0 L of solution?
 What weight of sodium chloride is needed for a volume of 2.0 L of solution?
 If needed, write out the conversion factors needed and include a final answer.
2. A 20.0 g sample of potassium chloride is dissolved in enough water to make 200 mL of solution. Calculate the mass percent of the potassium chloride.

$$\text{NaCl}_{\text{aq}} = \frac{1 \text{ mol}}{58.44 \text{ g}} \times 1.5 \text{ mol/L} \times 1.0 \text{ L} = 0.258 \text{ mol}$$

$$\text{NaCl}_{\text{aq}} = \frac{1 \text{ mol}}{58.44 \text{ g}} \times 1.5 \text{ mol/L} \times 2.0 \text{ L} = 0.516 \text{ mol}$$

3. Find the molarity of a solution made from 200 g of CaCl_2 dissolved in enough water to make 0.5 L.

$$\text{CaCl}_{2\text{aq}} = \frac{1 \text{ mol}}{110.98 \text{ g}} \times 200 \text{ g} = 1.80 \text{ mol}$$

$$\text{CaCl}_{2\text{aq}} = \frac{1 \text{ mol}}{110.98 \text{ g}} \times 200 \text{ g} = 1.80 \text{ mol}$$

4. An unknown ionic solution's "molarity" is defined by preparing 1.0 L of solution containing 1.0 g of unknown compound. Calculate the molarity of the solution.

$$\text{NaCl}_{\text{aq}} = \frac{1 \text{ mol}}{58.44 \text{ g}} \times 1.0 \text{ g} = 0.017 \text{ mol}$$

$$\text{NaCl}_{\text{aq}} = \frac{1 \text{ mol}}{58.44 \text{ g}} \times 1.0 \text{ g} = 0.017 \text{ mol}$$

5. A 20.0 g sample of sodium chloride is dissolved in water to produce a 0.5 L solution. Calculate the molarity of the solution.

$$\text{NaCl}_{\text{aq}} = \frac{1 \text{ mol}}{58.44 \text{ g}} \times 20.0 \text{ g} = 0.342 \text{ mol}$$

Unit 11 Worksheet 3 Molar Concentration

SJ Ball



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