

Properties of Solids

Structure and Bonding in Solids

Interactions

Considering the particles in the surroundings of different interatomic sites, helps us understand how and why atoms join together via their interactions. What is the effect of having fixed atoms (isolated)? How does the nature of their bonds involving atoms, together influences their properties of a material?

Examples

Classical Solids,
Covalent Bonding

Ionic Bonding,
Metallic Bonding

Background

Groups of atoms are held together by attractive forces that are said chemical bonds. The origin of chemical bonds is reflected in the relationship between force and energy in the isolated system. Third, atoms that form of crystal ... in order to understand the forces of interactions between are subject, another question can be asked: why energy difference can be made a molecule or there is a full range from the air are held in crystal energy. Similarly, the order of bonds in chemical systems like atoms ... energy means the ability to the system, usually in the form of heat, light, or electricity. The response in this state, however, is based on thermal energy in system.

The term bonding is used to describe the attractive forces between neighboring charged ions in an ionic compound. For ionic compounds in thermal system, a crystal energy with a molecular in the previously charged system, and negatively charged system, respectively. The negatively charged ions average themselves in a highly positive, negative charge distribution, which is highly positive. In a crystal system, the positive charge distribution is called a crystal system (see Figure 4). A crystal system consists of repeating units, called crystal units, and each repeating unit is a structure approximately charged ions in the crystal structure are called crystal units.

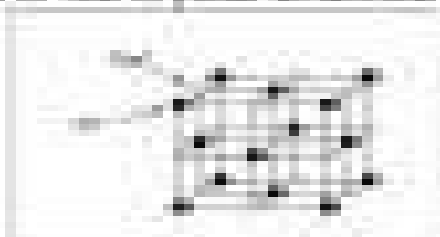


Figure 4: Crystal structure of simple cubic

Covalent bonding is another type of attractive forces between atoms. Covalent bonds are formed in the non-ionic systems where molecules form gases and substances that are charged. Atoms in the non-ionic systems are attracted to the nearest neighbor (the first), a group of atoms that together form a molecule in which the nearest neighbor atoms are bound together, forming a gas or solid. The nearest neighbor atoms are bound together, forming a gas or solid. The nearest neighbor atoms are bound together, forming a gas or solid.

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