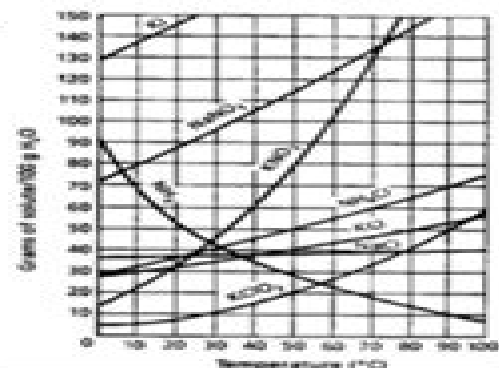


Name \_\_\_\_\_  
 Chemistry \_\_\_\_\_ Period \_\_\_\_\_  
 Date \_\_\_\_\_

## SOLUTIONS WORKSHEET

A. Solubility curves – refer to the graph to answer the following questions.



1. What is the solubility of  $\text{NH}_4\text{Cl}$  at  $70^\circ\text{C}$ ? **60 g**
2. Calculate the mass of  $\text{NH}_4\text{Cl}$  that would dissolve in 200 g of water at  $90^\circ\text{C}$ .

**70 g @ 100g water  $\rightarrow$  140 g**

3. At  $0^\circ\text{C}$ , 50.0 g  $\text{NH}_4\text{Cl}$  are added to 100 g water. To what temperature does the solution have to be raised in order to achieve a saturated solution of  $\text{NH}_4\text{Cl}_{\text{(aq)}}$ ?

**$50^\circ\text{C}$**

4. Is  $\text{NH}_3$  a (s), (l), or (g)? Explain.

**GAS – it's solubility decreases with increase in temperature. Opposite for solids**

5. Determine how much water is required to dissolve 35.0 g  $\text{NH}_4\text{Cl}$  at  $70^\circ\text{C}$  in order to make a saturated solution. Show your work.

$$\frac{60}{100} = \frac{35}{x}$$

$$x = 58.3$$

$$x = 58.3 \text{ g H}_2\text{O}$$

6. If a saturated solution of  $\text{NH}_3$  at  $15^\circ\text{C}$  is warmed to  $40^\circ\text{C}$ , how much ammonia will "fall out" of the solution?

**@  $15^\circ\text{C}$  – 60 g soluble**

**@  $40^\circ\text{C}$  – 35 g soluble**

**25 g will fall out**

## B. MOLARITY

1. What is molarity a measure of? What is the formula for molarity? What are the units of molarity?

**Concentration.**

$$M = n / V$$

**mol/L or M**

2. Calculate the molarity of a 1.5 L solution that has 0.45 mol solute dissolved.

**0.30 - M**

3. Calculate the grams of sodium sulfate that are required to prepare 1.50 L of 0.25 M solution.

**53.3 g  $\text{Na}_2\text{SO}_4$**

4. Calculate the volume of solution required to make 0.024 M solution using 15.0 g of calcium chloride.

**V = 5.6 L**

5. Describe in detail the steps you would take to prepare 1.50 L of 0.5 M solution of sodium chloride.

$$n = 0.75 \text{ mol} \rightarrow 43.5 \text{ g NaCl}$$

1) measure out 43.5 g NaCl

2) put NaCl into graduated cylinder

3) add enough water to make final volume 1.50 L

# Properties Of Solution Chemistry Workbook Answer Key

**Lingjun Ying**



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