## CLASSIFYING CHEMICAL REACTIONS



Classify the following reactions as synthesis, decomposition, single replacement or double replacement. De composition 2KCIO, → 2KCI + 3O, Double Replacement HCI + NaOH → NaCI + H<sub>\*</sub>O Single Replacement. Mg + 2HCl → MgCl, + H, Synthesis 2H<sub>2</sub> + O<sub>2</sub> → 2H<sub>2</sub>O Single Replacement 2AI + 3NIBr<sub>2</sub> → 2AIBr<sub>3</sub> + 3NI Synthesis 4AI + 3O<sub>5</sub> → 2AI<sub>2</sub>O<sub>5</sub> Decomposition 2NaCl → 2Na + Cl<sub>a</sub> Single Replacement CaCl<sub>2</sub> + F<sub>2</sub> → CaF<sub>2</sub> + Cl<sub>2</sub> Double Replacement AgNO<sub>3</sub> + KCI → AgCI + KNO<sub>3</sub> Synthesis N<sub>2</sub> + 3H<sub>2</sub> → 2NH<sub>3</sub> Decomposition 2H<sub>2</sub>O<sub>3</sub> → 2H<sub>2</sub>O + O<sub>3</sub> 12. (NH,)250, + Ba(NO3)2 - Baso, + 2NH,NO, Double Replacement Single Replacement 13. Mgl₂ + Br₂ → MgBr₂ + l₂ Synthesis 14. SO, + H₂O → H₂SO₄

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**E Durkheim** 

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