

# 12 Stoichiometry



## THE MOLE AND QUANTIFYING MATTER, REACTIONS

### 12.1 The Arithmetic of Equations

For students using the Foundation edition, assign problems 2–4, 7, 10–12.

#### Essential Understanding

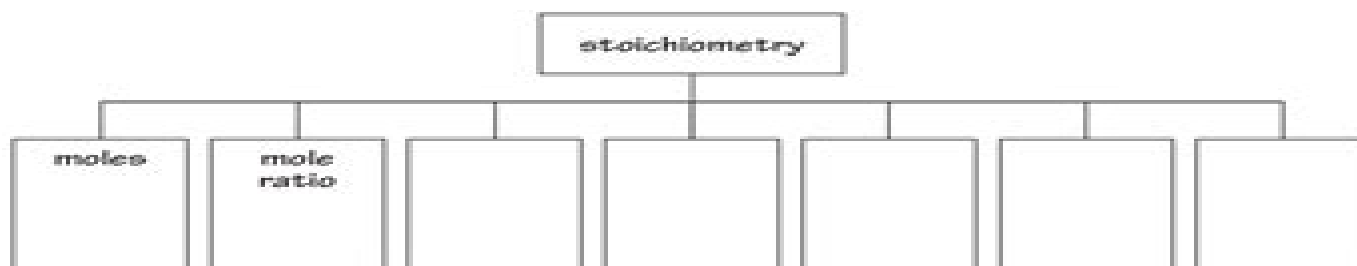
The law of conservation of mass applies to all chemical equations.



#### Reading Strategy

**Vocabulary Word Map** A vocabulary word map will help you learn vocabulary by associating the word with related words and images. Begin by writing the word *stoichiometry* in the top box.

As you read Lesson 12.1, use the word map below to help you get a better understanding of the meaning of the word *stoichiometry*. As you read, fill in the other boxes with terms, phrases, or images that are associated with the word.



#### Lesson Summary

**Using Equations** Stoichiometric calculations tell us the amounts of reactants and products under ideal conditions.

- ▶ Knowing the quantity of one substance in an equation allows you to calculate the amount of any other substance consumed or created in the reaction.

**Chemical Equations** The solution to every stoichiometric problem requires a balanced chemical equation.

- ▶ In a balanced chemical equation, the total number of atoms of each element in the reactants must equal the total number of atoms of that element in the products.
- ▶ In a balanced chemical equation, the total mass of the reactants must equal the total mass of the products.

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