Exam #10 Review: Acids, Bases, and pH

Know the properties of acids and bases. What are their similarities and differences?

Acids

- Taste sour
- React with metals to form hydrogen gas
- Contain a hydrogen ion (H+1)
- pH value less than 7

Bases

- Taste bitter
- Feel slippery
- Contain a hydroxide ion (OH-1)
- pH value greater than 7

BOTH

- · change colors of indicators
- · react with each other to form salt and water
- conduct electricity when dissolved in solution (electrolytes)
- Know the rules for naming acids (binary acid and oxyacid) and name the following below:

Binary acids (only 2 elements): hydro – element – ic acid Oxyacids (contain a polyatomic ion): NO "HYDRO" IN THE NAME

-ate ion → -ic acid
 -ite ion → -ous acid

- a. HCl hydrochloric acid b. HF hydrofluoric acid
- c. HClO₂ chlorous acid

REMEMBER:
"I ate something icky,
all nite I was nauseous"

- d. HNO₃ nitric acid e. H₃PO₄ phosphoric acid
- f. H₂SO₃ sulfurous acid
- Know the rules for naming bases and name the following below:

Element + hydroxide

- a. NaOH sodium hydroxide
- b. LiOH lithium hydroxide

- c. Al(OH)₃ aluminum hydroxide
- d. Ba(OH)₂ barium hydroxide

4. How are Arrehenius acids and bases defined?

Acids – produce H⁺¹ ions in aqueous solutions

Bases - produce OH-1 ions in aqueous solutions

5. How are Bronsted-Lowry acids and bases defined?

Acids – donate (give away) a H*1 ion (proton)
Bases – accept (take in) a H*1 ion (proton)

- 6. Show how the following acids and bases ionize:
 - a. HCl → H*1 + Cl-1
 - b. NaOH → Na+1 + OH-1

- c. $H_2CO_3 \rightarrow 2H^{+1} + CO_3^{-2}$
- d. Ba(OH)₂ → Ba*² + 2OH⁻¹
- If it is a strong acid or base, how much do they ionize in solution?STRONG acids & bases COMPLETELY ionize (all break apart) in a solution
- If it is a weak acid or base, how much do they ionize in solution?
 WEAK acids & bases do NOT completely ionize
- What is a neutralization reaction? What products are always produced?
 Acid reacts with a base always produces water and a salt
- 10. Complete the following neutralization reaction and what type of salt is made:

SA H₂SO₄ + KOH → H₂O + K₂SO₄

Ionization: $2H^{+1} + 5O_4^{-2} + K^{+1} + OH^{-1}$

Type of Salt: ____neutral____

SA SB HBr + Sr(OH)₂ →

Ionization: H+1 + Br1 + Sr+1 + 20H-1

H₂O + SrBr₂

Type of Salt: neutral

Worksheet 10 1 Acid Base Ph

Michael Clugston, Rosalind Flemming

Worksheet 10 1 Acid Base Ph:

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