

Chapter 13: Properties of Solutions

Answers to selected problems:

23. a) unsaturated b) saturated

24. a) 25 g

33. Solubility of Helium: 5.6×10^{-4} M; Solubility of N_2 = 9.0×10^{-4} M

34. Solubility of Oxygen: 2.5×10^{-4} M. Hint: Be sure to change torr to atm.

35. a) 2.15 %; b) 3.15 ppm

38. a) Mole fraction of phenol: 0.0285

b) 5.66 % c) 0.638 m

40. a) 0.125 M Aluminum Sulfate b) 0.140 M

c) 0.63 M (hint: use $MV = MV$, see Chapter 4)

41. a) 4.70 m b) 0.235 m

43. **To make this problem easier, I would assume 1.000 L or 1000 mL.

a) 43.01% b) mole fraction = 0.122

c) 7.69 m d) 5.828 M

46. a) 0.0439 b) 0.498 m c) 0.417 M

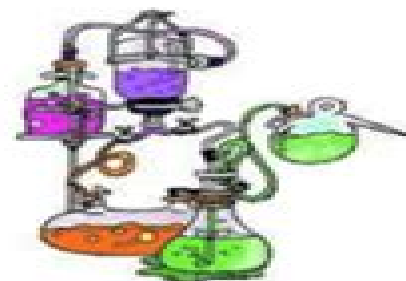
(Hint: on this problem, you will have to use the density of thiophene to get the mL of thiophene and you will also have to use the density of toluene to get the mass(g) of toluene.

48. a) 0.278 mol b) 6.25×10^{-5} moles (assume the mass of the solvent is approx. equal to the mass of the solution for this problem)

c) 3.29×10^{-3} mol

51. ** Assume you have 1.000L or 1000 mL. You can also use density to get the total grams.

Answer: 71%



Properties Of Solutions Guided Answers

B Lingard



Properties Of Solutions Guided Answers:

Unveiling the Power of Verbal Artistry: An Mental Sojourn through **Properties Of Solutions Guided Answers**

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Table of Contents Properties Of Solutions Guided Answers

1. Understanding the eBook Properties Of Solutions Guided Answers
 - The Rise of Digital Reading Properties Of Solutions Guided Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Properties Of Solutions Guided Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Properties Of Solutions Guided Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Properties Of Solutions Guided Answers
 - Personalized Recommendations

- Properties Of Solutions Guided Answers User Reviews and Ratings
- Properties Of Solutions Guided Answers and Bestseller Lists
- 5. Accessing Properties Of Solutions Guided Answers Free and Paid eBooks
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 - Highlighting and Note-Taking Properties Of Solutions Guided Answers
 - Interactive Elements Properties Of Solutions Guided Answers
- 8. Staying Engaged with Properties Of Solutions Guided Answers
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 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Properties Of Solutions Guided Answers
- 9. Balancing eBooks and Physical Books Properties Of Solutions Guided Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Properties Of Solutions Guided Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Properties Of Solutions Guided Answers
 - Setting Reading Goals Properties Of Solutions Guided Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Properties Of Solutions Guided Answers
 - Fact-Checking eBook Content of Properties Of Solutions Guided Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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