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Date: 8/4/20

Qualitative Analysis of Anions

Purpose of the Experiment

This lab introduces qualitative analysis, the area of chemistry concerned with the identification of substances by their chemical reactions. You will observe the reactions of some simple salts, and identify unknown samples. This experiment also provides a review of acid-base chemistry, introduces oxidation-reduction reaction and complex ion formation.

Introduction

Identifying samples of unknown substances is an important part of chemistry, with applications in medicine, environmental science, and geology. Materials can be characterized by both observing their characteristic chemical reactions, and by using instrumental analysis. Often identification of substances by their typical reactions is simple, accurate, and inexpensive, in comparison to instrumental methods. In this lab you will observe some characteristic reactions of six negatively charged ions (anions). Your observations of carbonate (CO_3^{2-}), bicarbonate (HCO_3^-), sulfite (SO_3^{2-}), sulfate (SO_4^{2-}), chloride (Cl^-), and iodide (I^-) reactions will then help you to identify unknown samples.

Anion Reactions

The chemical reactions that you will observe in the lab videos will allow you to witness the behavior of salts of known composition with a number of reagents. Then you will determine which anion is present in each of the unknown samples assigned to you. For this lab, the anions are separated into three main classification groups:

Group 1 Affected by addition of dilute hydrochloric acid	Group 2 Affected by the addition of concentrated sulfuric ² acid	Group 3 Precipitate with barium chloride
Carbonate (CO_3^{2-})	Chloride (Cl^-)	Sulfate (SO_4^{2-})
Bicarbonate (HCO_3^-)	Bromide (Br^-)	Phosphate (PO_4^{3-})
Sulfite (SO_3^{2-})	Iodide (I^-)	
Sulfide (S^{2-})	Nitrate (NO_3^-)	
Nitrite (NO_2^-)		
Thiosulfate (SCN^-)		

Group 1 anions are all affected by the addition of dilute hydrochloric acid since they are all anions of weak acids. The Group 2 anions result in the formation of volatile products when a strong acid is added, and the Group 3 anions result in the formation of insoluble barium salts upon the addition of BaCl_2 . There will be additional confirmation experiments for each of these anions as well.

² The British English spelling of sulfuric acid is sulphuric acid; similarly sulfate can also be written as sulphate; sulfite can be sulphite; and thiosulfate as thiosulphate. The IUPAC spelling is sulfate, sulfuric acid, etc.

Qualitative Analysis Of Anions Lab Report

Gary S. Thorpe

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