

Chapter 10 Chemical Quantities Review

Part 1: Vocabulary

Label the following descriptions with the correct vocabulary word.

- | | |
|---|---|
| _____ mass | _____ percent composition. |
| _____ empirical formula | _____ empirical formula |
| _____ molar mass | _____ standard temperature and pressure (STP, 1 atm, 0°C) |
| _____ 10 ²³ | |
| _____ is a description of the relative amounts of each element in a compound _____ percent composition | |
| _____ the lowest whole-number ratio of the atoms of the elements in a compound _____ empirical formula | |
| _____ the volume of 1 mole of any gas at STP _____ 22.4 L | |
| _____ the smallest particle of an element or compound _____ molar mass | |
| _____ the number representing 6.02 x 10 ²³ representative particles of a substance _____ Avogadro's number | |
| _____ 1 mole = 6.02 x 10 ²³ particles _____ molar mass | |
| _____ the number representing the relative mass of an element or compound _____ 12.01 g/mol | |
| _____ the number representing the relative mass of an element _____ 12.01 g/mol | |

Part 2: Section Concept Review

10.1 The Mole: A Measurement

Complete the following statement to demonstrate knowledge of the terms and your understanding of the concepts involved with the section of the textbook corresponding to this section (don't give an answer).

Knowing what a mole represents, its mass, and its volume for a single quantity called the _____ mole _____ Representative representative particles for 1 mole of a substance is _____ Avogadro's number _____

For the mass of 1 mole of a compound, scientists use together the _____ molar mass _____ of the atoms making the compound. When we calculate the molar mass for any substance, the _____ molar mass _____ of the compound. There are _____ 6.02 x 10²³ _____ representative particles in 1 mole of any substance.

True or False

1 mole of each of these compounds has a mass of 1 mole.

True or False: 1 mole of chlorine contains 6.02 x 10²³ atoms.

True or False: The representative particles of sodium chloride is the molecule. **False**

True or False: 1 mole of H₂O is composed of one mole of hydrogen and two atoms of oxygen.

True or False: The molar mass of nitrogen gas (N₂) is 28.02 g/mol. **True**

Problems

Workbook Chapter 10 Chemical Quantities Answer Key

Tobias Bleicker

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Workbook Chapter 10 Chemical Quantities Answer Key:

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