

/5 K

/5 A

/6 T

/4 C

Attendance # _____

Part A: Multiple Choice

Circle the best answer

1. Two reactants are mixed in a bomb calorimeter and a reaction occurs. The temperature of the calorimeter drops. Which of the following statements regarding the reaction is true?

- a) The system is releasing heat.
b) The system is absorbing heat.
 c) The heat of reaction is equal to $q\Delta T$.
 d) The reaction is exothermic.
 e) A and D are correct.

2. Which of the following are exothermic?

- I. boiling water II. freezing water
 III. condensing steam IV. melting ice

- a) I and II b) I and IV **c) II and III**
 d) II and IV e) II, III, and IV

3. Which of the following will show a negative entropy change in the forward direction?

- I. $\text{SO}_{2(g)} + \text{CaO}_{(s)} \rightarrow \text{CaSO}_{3(s)}$
 II. $\text{Ni}_{(s)} + 2\text{HCl}_{(aq)} \rightarrow \text{H}_{2(g)} + \text{NiCl}_{2(aq)}$
 III. hydrogen and oxygen combine to form water molecules
 IV. air is pumped into a tire
 V. sugar dissolves in water

- a) I, III and IV** b) II and V c) II and IV
 d) III and V e) V only

Part B: Application

Complete the questions in the spaces provided. Be sure to record your answers with proper significant digits and units: (5 marks)

1. The reduction of iron (II) oxide (FeO) to metallic iron (Fe) can be performed using methane (CH_4) according to the following reaction:



Given the following heats of formation, determine the heat required to produce exactly 250.g of iron. (5 marks)

$$\Delta H^\circ_f \text{FeO}_{(s)} = -238 \text{ kJ/mol}$$

$$\Delta H^\circ_f \text{CH}_{4(g)} = -73.61 \text{ kJ/mol}$$

$$\Delta H^\circ_f \text{CO}_{2(g)} = -393.5 \text{ kJ/mol}$$

$$\Delta H^\circ_f \text{H}_2\text{O}_{(g)} = -241.8 \text{ kJ/mol}$$

$$\Delta H = [4 \times 0 \text{ kJ/mol} + 1 \times -393.5 \text{ kJ/mol} + 2 \times -241.8 \text{ kJ/mol}] - [1 \times -73.61 \text{ kJ/mol} + 4 \times -238 \text{ kJ/mol}]$$

$$\Delta H = [-877.1 \text{ kJ/mol}] - [-1025.61 \text{ kJ/mol}]$$

$$\Delta H = 149 \text{ kJ/mol CH}_4 \rightarrow 37.25 \text{ kJ/mol Fe}$$

$$n = m/M$$

$$= 250 \text{ g} / 55.845 \text{ g/mol}$$

$$= 4.477 \text{ mol}$$

$$37.25 \text{ kJ/mol} \times 4.477 \text{ mol} = 167 \text{ kJ}$$

4. Which of the following is the correct thermochemical equation for the formation of $\text{C}_6\text{H}_5\text{OH}_{(l)}$, which absorbs 235.2 kJ/mol when formed?

- a) $2\text{C}_{(s)} + 5/2 \text{H}_{2(g)} + 1/2 \text{O}_{2(g)} \rightarrow \text{C}_6\text{H}_5\text{OH}_{(l)} + 235.2 \text{ kJ}$
 b) $2\text{C}_{(s)} + 3 \text{H}_{2(g)} + 1/2 \text{O}_{2(g)} \rightarrow \text{C}_6\text{H}_5\text{OH}_{(l)} + 235.2 \text{ kJ}$
 c) $2\text{C}_{(s)} + 5/2 \text{H}_{2(g)} + 1/2 \text{O}_{2(g)} + 235.2 \text{ kJ} \rightarrow \text{C}_6\text{H}_5\text{OH}_{(l)}$
d) $2\text{C}_{(s)} + 3 \text{H}_{2(g)} + 1/2 \text{O}_{2(g)} + 235.2 \text{ kJ} \rightarrow \text{C}_6\text{H}_5\text{OH}_{(l)}$
 e) $\text{C}_6\text{H}_5\text{OH}_{(l)} + 3\text{O}_{2(g)} \rightarrow 2\text{CO}_{2(g)} + 3\text{H}_2\text{O}_{(l)} + 235.2 \text{ kJ}$

5. When solid sodium hydroxide (NaOH) is added to water, the beaker feels warm to the touch. Which statement(s) best describe(s) this observation?

- I. $\text{NaOH}_{(s)} \rightarrow \text{NaOH}_{(aq)} + 43.6 \text{ kJ}$**
 II. $\text{NaOH}_{(s)} \rightarrow \text{NaOH}_{(aq)} \Delta H = +43.6 \text{ kJ}$
 III. The reaction is endothermic
 IV. Heat is absorbed by the system, so it feels warmer
 V. Heat is absorbed by the reactants in order to decrease the entropy of the system.

- a) I only** b) II & III c) II, III and IV
 d) II and V e) V only

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