

Thermochemistry, Entropy and Reaction Rates Review (answers in red)

Thermochemistry pp. 214-215, 227, 236-238

You must be able to:

- explain the differences between an endothermic and exothermic reaction
- explain what the ΔH for a reaction tells you about the reaction
- apply the information contained in a thermochemical equation to other problems
- convert between calories, kilocalories, joules, and kilojoules
- solve a lab-type thermochemistry problem

Example 1: When ammonium dichromate is ignited, it decomposes and a great deal of heat is given off. If 2.8 g of ammonium dichromate were decomposed in a copper container weighing 0.45 kg, the heat given off by the reaction is completely absorbed by the copper container, and the copper's temperature increases from 18°C to 372°C, what would be the ΔH in kilojoules for the following equation? [The specific heat of copper is 0.092 calories/g °C]



calculate ΔH from standard heats of formation

Example 2: Calculate the ΔH for the reaction between aluminum and copper II chloride using standard heats of formation. (See thermodynamics table provided to you—treat the compounds as solids instead of aqueous.) You need a balanced equation to solve this.
 $\Delta H = -748.1 \text{ kJ}$

Example 3: Using the ΔH from example 2, calculate the amount of heat that would be absorbed or released by the reaction of 10 grams of aluminum with an excess of copper II chloride. 138.5 kJ

Entropy pp. 864-876

- define entropy
- explain the effects endothermic/exothermic reactions have on the entropy of the surroundings
- determine how the entropy of a system changes given its chemical equation

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