

Thermochemistry Overview

Vocabulary

Thermochemistry: the study of energy changes during a chemical reaction

Energy Types

Kinetic Energy

- Energy associated with the movement of chemical species (particles)
- Moving electrons, vibration of atoms connected by chemical bonds, rotation and translation of molecules
- Related to temperature

Chemical Potential Energy

- Energy stored in covalent and/or ionic bonds of chemical reactants and products, or intermolecular forces (IMFs) like LDF or Dipole-Dipole forces

Enthalpy (ΔE)

- Refers to the energy change of the whole system
- The total of kinetic and potential energies within a chemical system
- Mostly considered to be potential energy

$$\Delta E_{\text{system}} = -\Delta E_{\text{surroundings}}$$

$$\Delta E_{\text{lost}} = -\Delta E_{\text{gained}}$$

Enthalpy Change (ΔH)

- The difference between the enthalpy of products and reactants
- Assume a difference in chemical potential only
- Also known as enthalpy of reaction
- Measured in kJ or J

$$\Delta H = \Delta H_{\text{products}} - \Delta H_{\text{reactants}}$$

$$\Delta H = n(H_m)$$

Molar Enthalpy (H_m OR H)

- The energy change **per mole** of one of the substances in a specific reaction
- Values are different for each substance, and different for each reaction
- Also known as molar enthalpy of reaction
- Measured in kJ/mol or J/mol

Thermochemistry Chemical Reaction Review

CL Gary



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