

Thermodynamic Feasible Reactions: (2000-2017)

10. Calculate the standard enthalpy of formation for the reaction



with $\Delta_f H^\circ$



with $\Delta_f H^\circ$ of BaSO_4 is

with $\Delta_f H^\circ$ of HClO_4 is

with $\Delta_f H^\circ$ of BaHClO_4 is

(2000-2017)

11. Calculate the free enthalpy of the combustion of vinyl acetate, with



with $\Delta_f H^\circ$ of CO_2 is



with $\Delta_f H^\circ$ of H_2O is



with $\Delta_f H^\circ$ of $\text{C}_4\text{H}_6\text{O}_2$ is

(2000-2017)

12. Calculate the total required for thermodynamic stability, CaCO_3 , relative CaO , and natural elements, with all natural elements at standard conditions



(2000-2017)

13. $\text{C}_2\text{H}_2 + 2\text{O}_2 \rightarrow 2\text{CO}_2 + \text{H}_2\text{O}$ $\Delta_f H^\circ = 0$



with $\Delta_f H^\circ$ of CO_2 is



with $\Delta_f H^\circ$ of H_2O is



(2000-2017)

14. Organometallics are the subject used that is sometimes used as a solvent. Calculate the standard enthalpy of formation for organometallics, given the following standard enthalpies of reaction



with $\Delta_f H^\circ$ of $\text{C}_2\text{H}_5\text{MgOH}$ is



with $\Delta_f H^\circ$ of $\text{C}_2\text{H}_5\text{MgOH}$ is



with $\Delta_f H^\circ$ of $\text{C}_2\text{H}_5\text{MgOH}$ is

(2000-2017)

15. The thermodynamic stability change is related to the thermodynamic of systems



Calculate the standard free enthalpy of formation for organometallics, the following standard enthalpies of formation: $\Delta_f H^\circ_{\text{BaO}} = -635 \text{ kJ mol}^{-1}$ with $\Delta_f H^\circ_{\text{BaSO}_4} = -1075 \text{ kJ mol}^{-1}$

(2000-2017)

16. The thermodynamic $\text{BaSO}_4 + \text{HClO}_4 \rightarrow \text{BaHClO}_4 + \text{H}_2\text{O}$

If you begin with 10.0 g of BaSO_4 , calculate the mass of BaHClO_4 that you would expect to produce

(2000-2017)

Thermochemistry Practice Problems Key

Raffaella Di Napoli



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