

## Thermochemistry Review - Answers

1. Using the data below, calculate the amount of energy required to melt a 16.2 g aluminum pop can. Assume that the aluminum is at an initial temperature of 20°C.

- Molar heat of fusion of Al(s) = 10.7 kJ/mol
- Melting point of Al(s) = 660°C
- Heat capacity of Al(s) = 0.900 J/g°C

$$\text{mol} = 16.2 \text{ g} \div 26.98154 \text{ g/mol} = 0.600 \text{ mol}$$

$$\Delta H = (20^\circ\text{C} \rightarrow 660^\circ\text{C}) + (\text{melt})$$

$$\Delta H = (mc\Delta T) + (n\Delta H_{\text{fus}})$$

$$\Delta H = (16.2 \text{ g})(0.900 \text{ J/g}^\circ\text{C})(640^\circ\text{C}) + (0.600 \text{ mol})(10.7 \text{ kJ/mol})$$

$$\Delta H = 9331.2 \text{ J} + 6.42 \text{ kJ}$$

$$\Delta H = 9.33 \text{ kJ} + 6.42 \text{ kJ}$$

$$\Delta H = \mathbf{15.8 \text{ kJ}}$$

2. Given the thermochemical equations:



Calculate the  $\Delta H$  for the reaction:  $\text{CH}_3\text{OH}(\text{l}) + \frac{3}{2} \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l})$



# Thermochemistry Review And Answers

**Sebastian Brünink**



## **Thermochemistry Review And Answers:**

This book delves into Thermochemistry Review And Answers. Thermochemistry Review And Answers is a crucial topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Thermochemistry Review And Answers, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Thermochemistry Review And Answers
- Chapter 2: Essential Elements of Thermochemistry Review And Answers
- Chapter 3: Thermochemistry Review And Answers in Everyday Life
- Chapter 4: Thermochemistry Review And Answers in Specific Contexts
- Chapter 5: Conclusion

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