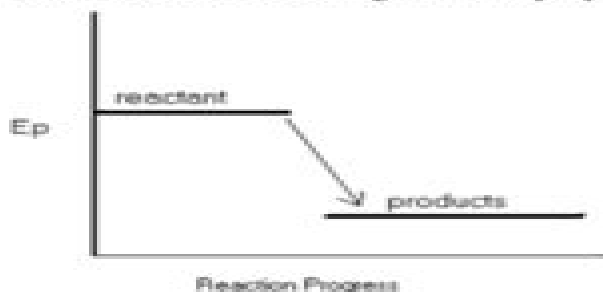


Thermodynamics and Rates of Reactions

Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

- _____ 1. An endothermic reaction is one where
- heat is transferred from the surroundings into a system
 - heat is transferred from a system into the surroundings
 - kinetic energy is transformed into potential energy
 - there is no transfer of heat
 - none of the above
- _____ 2. What is another way to write the following equation?
 $\text{C}_2\text{H}_4(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + 2\text{H}_2(\text{g}) \quad \Delta H = 3.2 \times 10^2 \text{ kJ/mol}$
- $\text{C}_2\text{H}_4(\text{g}) \rightarrow \text{C}_2\text{H}_4(\text{g}) + 2\text{H}_2(\text{g}) + 3.2 \times 10^2 \text{ kJ}$
 - $\text{C}_2\text{H}_4(\text{g}) + 3.2 \times 10^2 \text{ kJ} \rightarrow \text{C}_2\text{H}_2(\text{g}) + 2\text{H}_2(\text{g})$
 - $2\text{H}_2(\text{g}) + \text{C}_2\text{H}_2(\text{g}) + 3.2 \times 10^2 \text{ kJ} \rightarrow \text{C}_2\text{H}_4(\text{g})$
 - $\text{C}_2\text{H}_4(\text{g}) - 3.2 \times 10^2 \text{ kJ} \rightarrow \text{C}_2\text{H}_2(\text{g}) + 2\text{H}_2(\text{g})$
 - none of the above
- _____ 3. 1.00 mol of an element burned in a calorimeter raise the temperature of 100 g of water by 10.0°C . The molar heat of combustion of the element is
- 6.36 kJ/mol
 - 0.239 kJ/mol
 - 41.8 kJ/mol
 - 4.20 kJ/mol
 - 16.9 kJ/mol
- _____ 4. Which statement concerning the accompanying diagram is true?



- ΔH is positive
 - the system is endothermic
 - the system releases heat to the surroundings
 - the heat content of the reactants is less than the heat content of the products
 - the enthalpy of the products is greater than the enthalpy of reactants
- _____ 5. Which of the following equations represents the correct heat of formation of water vapour?
- $\text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$
 - $\text{H}_2(\text{g}) + \text{O}(\text{g}) \rightarrow \text{H}_2\text{O}(\text{g})$
 - $2\text{H}(\text{g}) + \text{O}(\text{g}) \rightarrow \text{H}_2\text{O}(\text{g})$
 - $\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{g})$
 - $\text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{g})$

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Christian Drosten



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