

## Chapter 4 Atomic Structure

**Section 4.1 Studying Atoms**

(pages 100–105)

*This section discusses the development of atomic models.***Reading Strategy (page 100)**

**Summarizing** As you read, complete the table about atomic models. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

Atomic Models		
Scientist	Evidence	Model
Dalton		Indivisible, solid spheres
	Deflected beam	Negative charges evenly scattered through positively charged mass of matter (plum pudding model)
	Deflection of alpha particles passing through gold foil	

**Ancient Greek Models of Atoms (page 100)**

- Democritus named the smallest particles of matter \_\_\_\_\_ because they could not be divided.

**Dalton's Atomic Theory (page 101)**

- Is the following sentence true or false? John Dalton gathered evidence for the existence of atoms by measuring the masses of elements that reacted to form compounds. \_\_\_\_\_
- Dalton's theory suggests that all matter is made up of individual particles called \_\_\_\_\_, which cannot be \_\_\_\_\_.
- Circle the letters of the sentences that represent main points of Dalton's theory of atoms.
  - All elements are composed of atoms.
  - In a particular compound, atoms of different elements always combine the same way.
  - All atoms have the same mass.

## Section 4 Studying Atoms Answer Key

**Robert Dalling**



## Section 4 Studying Atoms Answer Key:

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