

8.2

THE NATURE OF COVALENT BONDING

KEY

Section Review

Objectives

- State a rule that usually tells how many electrons are shared to form a covalent bond
- Describe how electron dot formulas are used
- Predict when two atoms are likely to be joined by a double or a triple covalent bond
- Distinguish between a single covalent bond and other covalent bonds
- Describe how the strength of a covalent bond is related to its bond dissociation energy
- Describe how resonance structures explain bonding
- Identify some exceptions to the octet rule

Vocabulary

- | | | |
|------------------------|----------------------------|----------------------------|
| • single covalent bond | • double covalent bonds | • polyatomic ion |
| • structural formulas | • triple covalent bonds | • bond dissociation energy |
| • unshared pairs | • coordinate covalent bond | • resonance structures |

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

When atoms share electrons to gain the 1 configuration of a noble gas, the bonds formed are 2. A 3 pair of valence electrons constitutes a 4 covalent bond. Pairs of valence electrons that are not shared between atoms are called 5. Sometimes two or three pairs of electrons may be shared to give 6 covalent bonds. In some cases, only one of the atoms in a bond provides the pair of bonding electrons; this is a 7. 8 is required to break covalent bonds between atoms. The total energy required to break the bond between two covalently bonded atoms is known as the 9.

When it is possible to write two or more valid electron dot formulas for a molecule or ion, each formula is referred to as a 10.

- stable
- covalent
- shared
- single
- unshared pairs
- double/triple
- coordinate covalent bond
- energy
- bond dissociation energy
- resonance structure

Section 8 2 The Nature Of Covalent Bonding Worksheet

**Judith Kinnear, Marjory Martin, Lucy
Cassar, Elise Meehan, Ritu Tyagi**



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Table of Contents Section 8 2 The Nature Of Covalent Bonding Worksheet

1. Understanding the eBook Section 8 2 The Nature Of Covalent Bonding Worksheet
 - The Rise of Digital Reading Section 8 2 The Nature Of Covalent Bonding Worksheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Section 8 2 The Nature Of Covalent Bonding Worksheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Section 8 2 The Nature Of Covalent Bonding Worksheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Section 8 2 The Nature Of Covalent Bonding Worksheet
 - Personalized Recommendations
 - Section 8 2 The Nature Of Covalent Bonding Worksheet User Reviews and Ratings
 - Section 8 2 The Nature Of Covalent Bonding Worksheet and Bestseller Lists
5. Accessing Section 8 2 The Nature Of Covalent Bonding Worksheet Free and Paid eBooks
 - Section 8 2 The Nature Of Covalent Bonding Worksheet Public Domain eBooks
 - Section 8 2 The Nature Of Covalent Bonding Worksheet eBook Subscription Services
 - Section 8 2 The Nature Of Covalent Bonding Worksheet Budget-Friendly Options

6. Navigating Section 8 2 The Nature Of Covalent Bonding Worksheet eBook Formats
 - ePub, PDF, MOBI, and More
 - Section 8 2 The Nature Of Covalent Bonding Worksheet Compatibility with Devices
 - Section 8 2 The Nature Of Covalent Bonding Worksheet Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Section 8 2 The Nature Of Covalent Bonding Worksheet
 - Highlighting and Note-Taking Section 8 2 The Nature Of Covalent Bonding Worksheet
 - Interactive Elements Section 8 2 The Nature Of Covalent Bonding Worksheet
8. Staying Engaged with Section 8 2 The Nature Of Covalent Bonding Worksheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Section 8 2 The Nature Of Covalent Bonding Worksheet
9. Balancing eBooks and Physical Books Section 8 2 The Nature Of Covalent Bonding Worksheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Section 8 2 The Nature Of Covalent Bonding Worksheet
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Section 8 2 The Nature Of Covalent Bonding Worksheet
 - Setting Reading Goals Section 8 2 The Nature Of Covalent Bonding Worksheet
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Section 8 2 The Nature Of Covalent Bonding Worksheet
 - Fact-Checking eBook Content of Section 8 2 The Nature Of Covalent Bonding Worksheet
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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type 4 it does not required more effort at turning

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