

## % YIELD CALCULATIONS

The relative formula masses of the compounds mentioned in the question are shown below the formulae of the compounds

1. Calculate the percentage yield of tungsten (W) in the following reaction where 565g of tungsten(VI) oxide ( $\text{WO}_3$ ) produces 385g of tungsten.



$$\text{100\% yield} = (184 / 232) \times 565 = 448.1 \text{ g} \quad (\checkmark)$$

$$\% \text{ yield} = (385 / 448.1) \times 100 = 85.9 \% \quad (\checkmark)$$

2. Calculate the percentage yield of titanium (Ti) in the following reaction where 260g of titanium(IV) chloride ( $\text{TiCl}_4$ ) produces 48.5g of titanium.



$$\text{100\% yield} = (48 / 190) \times 260 = 65.7 \text{ g} \quad (\checkmark)$$

$$\% \text{ yield} = (48.5 / 65.7) \times 100 = 73.8 \% \quad (\checkmark)$$

3. Calculate the percentage yield of sulfur dioxide ( $\text{SO}_2$ ) in the following reaction where 5g of zinc sulphide ( $\text{ZnS}$ ) produces 2.8g of sulfur dioxide.



$$\text{Mr of ZnS} = 97 \quad \text{Mr of SO}_2 = 64 \quad (\checkmark)$$

$$\text{100\% yield} = (64 / 97) \times 5 = 3.299 \text{ g} \quad (\checkmark)$$

$$\% \text{ yield} = (2.8 / 3.299) \times 100 = 84.9 \% \quad (\checkmark)$$

4. Calculate the percentage yield of ethene ( $\text{C}_2\text{H}_4$ ) in the following reaction where 28g of decane ( $\text{C}_{10}\text{H}_{22}$ ) produces 3.3g of ethene.



$$\text{Mr of C}_{10}\text{H}_{22} = 142 \quad \text{Mr of C}_2\text{H}_4 = 28 \quad (\checkmark)$$

$$\text{100\% yield} = (28 / 142) \times 28 = 5.52 \text{ g} \quad (\checkmark)$$

$$\% \text{ yield} = (3.3 / 5.52) \times 100 = 59.8 \% \quad (\checkmark)$$

# Section Quiz Limiting Reactants And Percentage Yield

**James C. Hill**



## Section Quiz Limiting Reactants And Percentage Yield:

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## **Table of Contents Section Quiz Limiting Reactants And Percentage Yield**

1. Understanding the eBook Section Quiz Limiting Reactants And Percentage Yield
  - The Rise of Digital Reading Section Quiz Limiting Reactants And Percentage Yield
  - Advantages of eBooks Over Traditional Books
2. Identifying Section Quiz Limiting Reactants And Percentage Yield
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Section Quiz Limiting Reactants And Percentage Yield
  - User-Friendly Interface
4. Exploring eBook Recommendations from Section Quiz Limiting Reactants And Percentage Yield
  - Personalized Recommendations
  - Section Quiz Limiting Reactants And Percentage Yield User Reviews and Ratings
  - Section Quiz Limiting Reactants And Percentage Yield and Bestseller Lists
5. Accessing Section Quiz Limiting Reactants And Percentage Yield Free and Paid eBooks
  - Section Quiz Limiting Reactants And Percentage Yield Public Domain eBooks

- Section Quiz Limiting Reactants And Percentage Yield eBook Subscription Services
- Section Quiz Limiting Reactants And Percentage Yield Budget-Friendly Options
- 6. Navigating Section Quiz Limiting Reactants And Percentage Yield eBook Formats
  - ePub, PDF, MOBI, and More
  - Section Quiz Limiting Reactants And Percentage Yield Compatibility with Devices
  - Section Quiz Limiting Reactants And Percentage Yield Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Section Quiz Limiting Reactants And Percentage Yield
  - Highlighting and Note-Taking Section Quiz Limiting Reactants And Percentage Yield
  - Interactive Elements Section Quiz Limiting Reactants And Percentage Yield
- 8. Staying Engaged with Section Quiz Limiting Reactants And Percentage Yield
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Section Quiz Limiting Reactants And Percentage Yield
- 9. Balancing eBooks and Physical Books Section Quiz Limiting Reactants And Percentage Yield
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Section Quiz Limiting Reactants And Percentage Yield
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Section Quiz Limiting Reactants And Percentage Yield
  - Setting Reading Goals Section Quiz Limiting Reactants And Percentage Yield
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Section Quiz Limiting Reactants And Percentage Yield
  - Fact-Checking eBook Content of Section Quiz Limiting Reactants And Percentage Yield
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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