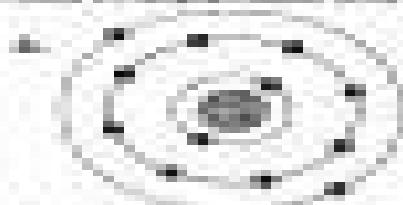


IPS Unit 9 – Chemical Bonds

Types of Bonds Worksheet

Directions: Study the diagrams below, then give a brief description of each bond type using the space provided.



3. Fill in the blanks (use arrows to show the flow).

- How many electrons will flow in each? _____
- How many electrons will stay in each? _____
- What will be the charge of atom A? _____
- What will be the charge of atom B? _____
- What will be the bond charge of the completed bond? _____
- What type of bond will form? _____

4. Explain why a covalent bond is stronger than ionic bonds (use the periodic table to check the charges).

State and show how electronegativity values of hydrogen, chlorine, and oxygen are used to predict the type of bond that will form between them.

Directions: Complete the table by comparing the properties of ionic and covalent compounds.

Compounds	Ionic compounds	Covalent compounds
How they are formed (electron transfer)	Electron transfer from one atom to another, resulting in the formation of ions.	Electron sharing between two atoms, resulting in the formation of a covalent bond.
Electron pairs	None	Shared pairs
How they are separated (heat, phase of matter, etc.)	High	Low

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