

CALCULATING REACTING QUANTITIES

Syllabus reference 8.3.4

For multiple-choice questions 1–5 circle the letter corresponding to the most correct answer.

- 1 The balanced equation for the burning of butanol ($\text{C}_4\text{H}_9\text{OH}$) is given below:



The amount of oxygen needed to burn 2 mole of butanol is:

- a 2 moles
 - b 4 moles
 - c 6 moles
 - ☒ d 12 moles
- 2 For the following reaction:
- $$\text{PbSO}_4(\text{s}) + \text{H}^+ \rightarrow \text{Pb}^{2+} + \text{HSO}_4^-$$
- Which of the following is NOT true?
- a One mole of lead sulfate produces one mole of lead ions.
 - b The reaction takes place in water.
 - ☒ c Both products precipitate out of solution.
 - d The equation is balanced.
- 3 A student has two identical flasks. One flask is filled with 1 mole of carbon dioxide (CO_2) at a temperature of 25°C and the other flask is filled with 2 moles of hydrogen (H_2) at a temperature of 0°C .
- The ratio of number of molecules of CO_2 to the number of molecules of H_2 is:
- a 4:1
 - b 2:1
 - c 1:1
 - ☒ d 1:2
- 4 The law of combining volumes was proposed by:
- a Dalton
 - ☒ b Gay-Lussac
 - c Avogadro
 - d Newton

Section Calculating Quantities In Reactions Answer Key

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Parsons, Gwen Pilling, Gareth Price**

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