#### SECTION 2-1

#### REVIEW AND REINFORCE

## Solids, Liquids, and Gases

#### Understanding Main Ideas

Use the diagram to answer Questions 1 through 3. Write your answers on a separate sheet of paper.

- Identify the physical state of the substances pictured.
- What would happen to the shape of each substance if the jars were broken? Use the differences in the physical state of the substances to explain your answer.



2

Would the volume of each substance change if each were moved into a larger container? Explain.

#### Building Vocabulary

Wirite a definition for each of the following terms in the spaces provided.

- a tixed position, definite shape & volume.
- 5. Equid particles are slightly less packed than solids,
- nowing out high speeds in all chrections, indefinite shape, indefinite valume
  - 7. VISCOSITY THE Slower the flowing. Wi the higher the
- 8. amorphous sold particles are not arranged in a regular pattern,
- forming crystals, definite melting front
- 10. must a Substance that flaux, & liquid or gos

# **Solids Liquids And Gases Review And Reinforce**

Saleh Al Arni, Sudip Chakraborty, Yeek Chia Ho, Woei Jye Lau, Rajamohan Natarajan

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web jan 29 2020 step 3 determine the number of bonds in the molecule covalent bonds are formed when one electron from each atom forms an electron pair step 2 tells how many electrons are needed and step 1 is how many electrons you have subtracting the number in step 1 from the number in step 2 gives you the number of electrons needed to complete **vsepr using lewis structures and vsepr to determine molecular geometry** - May 12 2023

web draw the lewis structure for each of the following compounds using vsepr determine the electron pair geometry the molecular geometry and the bond angle in degrees for each compound

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web this effect determines the atom s geometry and bond angles two regions will be 180 apart three regions will be 120 apart and four regions will be 109 5 apart geometry determination a determine the lewis dot structure of the molecule or ion b for each central atom in the structure determine the areas of electron density that lie

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