

Table 5-7

Solubility-Product Constants, K_{sp} , at 25°C

Fluorides		Chromates (cont.)		Hydroxides (cont.)			
BaF ₂	2.4 × 10 ⁻⁶	Ag ₂ CrO ₄	1.9 × 10 ⁻¹²	Ni(OH) ₂	1.6 × 10 ⁻¹⁶		
MgF ₂	8 × 10 ⁻⁸	PbCrO ₄	2 × 10 ⁻¹⁶	Zn(OH) ₂	4.5 × 10 ⁻¹⁷		
PbF ₂	4 × 10 ⁻⁸	Carbonates		Cu(OH) ₂	1.6 × 10 ⁻¹⁹		
SrF ₂	7.9 × 10 ⁻¹⁰			Hg(OH) ₂	3 × 10 ⁻²⁶		
CaF ₂	3.9 × 10 ⁻¹¹			Sr(OH) ₂	3 × 10 ⁻³⁷		
Chlorides				NiCO ₃	1.4 × 10 ⁻⁷	Cr(OH) ₃	6.7 × 10 ⁻³¹
				CaCO ₃	4.7 × 10 ⁻⁹	Al(OH) ₃	5 × 10 ⁻³³
		BaCO ₃	1.6 × 10 ⁻⁹	Fe(OH) ₃	6 × 10 ⁻³⁸		
		SrCO ₃	7 × 10 ⁻¹⁰	Co(OH) ₃	2.5 × 10 ⁻⁴³		
PbCl ₂	1.6 × 10 ⁻⁵	CuCO ₃	2.5 × 10 ⁻¹⁰	Sulfides			
AgCl	1.7 × 10 ⁻¹⁰	ZnCO ₃	2 × 10 ⁻¹⁰				
Hg ₂ Cl ₂ *	1.1 × 10 ⁻¹⁸	MnCO ₃	8.8 × 10 ⁻¹¹				
Bromides		FeCO ₃	2.1 × 10 ⁻¹¹				
		Ag ₂ CO ₃	8.2 × 10 ⁻¹²				
		CdCO ₃	5.2 × 10 ⁻¹²				
		PbCO ₃	1.5 × 10 ⁻¹³				
Iodides		MgCO ₃	1 × 10 ⁻¹⁵				
		Hg ₂ CO ₃ *	9.0 × 10 ⁻¹⁵				
		Hydroxides					
Sulfates		Ba(OH) ₂	5.0 × 10 ⁻³	Phosphates			
		Sr(OH) ₂	3.2 × 10 ⁻⁴				
		Ca(OH) ₂	1.3 × 10 ⁻⁶				
		AgOH	2.0 × 10 ⁻⁸				
		Mg(OH) ₂	8.9 × 10 ⁻¹²				
BaSO ₄	1.5 × 10 ⁻⁹	Mn(OH) ₂	2 × 10 ⁻¹³			Ag ₃ PO ₄	1.8 × 10 ⁻¹⁸
Chromates		Cd(OH) ₂	2.0 × 10 ⁻¹⁴			Sr ₃ (PO ₄) ₂	1 × 10 ⁻³¹
		Pb(OH) ₂	4.2 × 10 ⁻¹⁵			Ca ₃ (PO ₄) ₂	1.3 × 10 ⁻³²
		Fe(OH) ₂	1.8 × 10 ⁻¹⁵			Ba ₃ (PO ₄) ₂	6 × 10 ⁻³⁹
		Co(OH) ₂	2.5 × 10 ⁻¹⁶			Pb ₃ (PO ₄) ₂	1 × 10 ⁻⁵⁴
SrCrO ₄	3.6 × 10 ⁻⁵						
Hg ₂ CrO ₄ *	2 × 10 ⁻⁹						
BaCrO ₄	8.5 × 10 ⁻¹¹						

*As Hg₂²⁺ ion. $K_{sp} = [\text{Hg}_2^{2+}][\text{X}^-]^2$

Solubility Chart Chemistry 12

William M. Haynes

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Solubility Chart Chemistry 12:

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