

CHAPTER 10 The Mole

Section 1 Measuring Matter

pages 320–324

Practice Problems

pages 323–324

1. Zinc (Zn) is used to form a corrosion-inhibiting surface on galvanized steel. Determine the number of Zn atoms in 2.50 mol of Zn.

$$2.50 \text{ mol Zn} \times \frac{6.02 \times 10^{23} \text{ atoms Zn}}{1 \text{ mol Zn}} \\ = 1.51 \times 10^{24} \text{ atoms of Zn}$$

2. Calculate the number of molecules in 11.5 mol of water (H₂O).

$$11.5 \text{ mol H}_2\text{O} \times \frac{6.02 \times 10^{23} \text{ molecules H}_2\text{O}}{1 \text{ mol H}_2\text{O}} \\ = 6.92 \times 10^{24} \text{ molecules H}_2\text{O}$$

3. Silver nitrate (AgNO₃) is used to make several different silver halides used in photographic films. How many formula units of AgNO₃ are there in 3.25 mol AgNO₃?

$$3.25 \text{ mol AgNO}_3 \times \frac{6.02 \times 10^{23} \text{ formula units AgNO}_3}{1 \text{ mol AgNO}_3} \\ = 1.96 \times 10^{24} \text{ formula units AgNO}_3$$

4. **Challenge** Calculate the number of oxygen atoms in 5.00 mol of oxygen molecules. Oxygen is a diatomic molecule, O₂.

$$5.00 \text{ mol O}_2 \times \frac{6.02 \times 10^{23} \text{ molecules O}_2}{1 \text{ mol O}_2} \\ \times \frac{2 \text{ O atoms}}{\text{molecule O}_2} = 6.02 \times 10^{24} \text{ atoms O}$$

5. How many moles contain each of the following?
 - a. 5.75×10^{24} atoms Al

$$5.75 \times 10^{24} \text{ atoms Al} \times \frac{1 \text{ mol Al}}{6.02 \times 10^{23} \text{ atoms Al}} \\ = 9.55 \text{ mol Al}$$

- b. 2.50×10^{20} atoms Fe

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$$2.50 \times 10^{20} \text{ atoms Fe} \times \frac{1 \text{ mol Fe}}{6.02 \times 10^{23} \text{ atoms Fe}} \\ = 4.15 \times 10^{-4} \text{ mol Fe}$$

6. **Challenge** Identify the representative particle for each formula and convert the given number of representative particles to moles.

- a. 3.75×10^{24} CO₂

The representative particle is a molecule.

$$3.75 \times 10^{24} \text{ molecules CO}_2 \times \frac{1 \text{ mol CO}_2}{6.02 \times 10^{23} \text{ molecules CO}_2} = 6.23 \text{ mol CO}_2$$

- b. 3.58×10^{23} ZnCl₂

The representative particle is a formula unit.

$$3.58 \times 10^{23} \text{ formula units ZnCl}_2 \times \frac{1 \text{ mol ZnCl}_2}{6.02 \times 10^{23} \text{ formula units ZnCl}_2} \\ = 0.595 \text{ mol ZnCl}_2$$

Section 1 Review

page 324

7. **Explain** why chemists use the mole.

Chemists use the mole because it is a convenient way of knowing how many representative particles are in a sample.

8. **State** the mathematical relationship between Avogadro's number and 1 mol.

One mole contains Avogadro's number (6.02×10^{23}) of representative particles.

Solution Manual Chemistry The Mole

Didier Musso



Solution Manual Chemistry The Mole:

The book delves into Solution Manual Chemistry The Mole. Solution Manual Chemistry The Mole is an essential topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Solution Manual Chemistry The Mole, encompassing both the fundamentals and more intricate discussions.

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 - Chapter 1: Introduction to Solution Manual Chemistry The Mole
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 - Chapter 4: Solution Manual Chemistry The Mole in Specific Contexts
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The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Solution Manual Chemistry The Mole.

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