

- 2.6** A steel plate of 150 kg acceleration from rest with 3 m/s^2 for a period of 10s. What force is needed and what is the final velocity?

Solution:

Constant acceleration can be integrated to get velocity.

$$a = dV / dt \Rightarrow \int dV = \int a dt \Rightarrow \Delta V = a \Delta t = 3 \times 10 = 30 \text{ m/s}$$

$$V = 30 \text{ m/s}; \quad F = ma = 150 \times 3 = 2000 \text{ N}$$

- 2.7** A 15 kg steel container has 1.75 kilograms of liquid propane inside. A force of 2 kN acts acceleration this system. What is the acceleration?

Solution:

$$ma = \sum F \Rightarrow a = \sum F / m$$

$$m = m_{\text{steel}} + m_{\text{propane}} = 15 + 1.75 = 16.75 = 16.75 \text{ kg} \Rightarrow 162.167 \text{ kg}$$

$$a = 2000 / 162.167 = 12.3 \text{ m/s}^2$$

- 2.8** A rope hangs over a pulley with the two equally long ends down. One end and you attach a mass of 5 kg and on the other end you attach 10 kg. Assuming standard gravitation and no friction in the pulley, what is the acceleration of the 10 kg mass when released?

Solution:

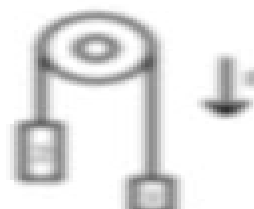
On the equation of motion for the mass m_1 along the downwards direction, in that case the mass m_1 moves up (i.e. has $-a$ for the acceleration)

$$m_1 a = m_1 g - m_1 g - m_1 a$$

$$(m_1 + m_1) a = (m_1 - m_1) g$$

This is not force in motion direction

$$a = (10 - 5)g / (10 + 5) = 5 \times g / 15 = 3.27 \text{ m/s}^2$$



- 2.9** A bucket of concrete of total mass 200 kg is raised by a crane with an acceleration of 2 m/s^2 relative to the ground at a location where the local gravitational acceleration is 9.5 m/s^2 . Find the required force.

Solution:

$$F = ma = F_{\text{up}} - mg$$

$$F_{\text{up}} = ma + mg = 200(2 + 9.5) = 2200 \text{ N}$$

Solution Manual For Fundamental Of Thermodynamics Van Wylen

Liyong Dong



Solution Manual For Fundamental Of Thermodynamics Van Wylen:

Chemical Engineering Thermodynamics AHUJA, PRADEEP, 2008-12 This book offers a full account of thermodynamic systems in chemical engineering It provides a solid understanding of the basic concepts of the laws of thermodynamics as well as their applications with a thorough discussion of phase and chemical reaction equilibria At the outset the text explains the various key terms of thermodynamics with suitable examples and then thoroughly deals with the virial and cubic equations of state by showing the P V T pressure molar volume and temperature relation of fluids It elaborates on the first and second laws of thermodynamics and their applications with the help of numerous engineering examples The text further discusses the concepts of exergy standard property changes of chemical reactions thermodynamic property relations and fugacity The book also includes detailed discussions on residual and excess properties of mixtures various activity coefficient models local composition models and group contribution methods In addition the text focuses on vapour liquid and other phase equilibrium calculations and analyzes chemical reaction equilibria and adiabatic reaction temperature for systems with complete and incomplete conversion of reactants Key Features Includes a large number of fully worked out examples to help students master the concepts discussed Provides well graded problems with answers at the end of each chapter to test and foster students conceptual understanding of the subject The total number of solved examples and end chapter exercises in the book are over 600 Contains chapter summaries that review the major concepts covered The book is primarily designed for the undergraduate students of chemical engineering and its related disciplines such as petroleum engineering and polymer engineering It can also be useful to professionals The Solution Manual containing the complete worked out solutions to chapter end exercises and problems is available for instructors

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