

3. Find the molarity of a solution made from 275 g of  $\text{CuSO}_4$  dissolved in enough water to make 4.25 L.

$$\frac{\text{mol}}{\text{dm}^3}$$

$$\frac{1.72 \text{ mol}}{4.25 \text{ dm}^3}$$

$$\boxed{\text{conc.} = \frac{.405 \text{ mol}}{\text{dm}^3}}$$

$$1 \text{ L} = 1 \text{ dm}^3$$

$$\text{Cu} = 63.5$$

$$\text{S} = 32.1$$

$$\text{O} \times 4 = \frac{64.0}{159.5 \text{ g/mol}}$$

$$\frac{159.5 \text{ g}}{1 \text{ mol}}$$

$$\frac{275 \text{ g}}{159.5 \text{ g}} = \frac{x \text{ mol}}{1 \text{ mol}}$$

$$x = 1.72 \text{ mol CuSO}_4$$

4. An alcoholic iodine solution ("tincture" of iodine) is prepared by dissolving 5.15 g of iodine crystals in enough alcohol to make a volume of 225 mL. Calculate the molarity of iodine in the solution.

$$\frac{5.15 \text{ g I}_2}{254 \text{ g I}_2} = \frac{x \text{ mol}}{1 \text{ mol}}$$

$$x = .0203 \text{ mol I}_2$$

$$x = .0203 \text{ mol I}_2$$

$$\frac{\text{mol}}{\text{dm}^3}$$

$$225 \text{ mL} = 225 \text{ cm}^3 =$$

# Solution Concentration Practice Problems

**Xiang Xie**



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