

ELECTRON ARRANGEMENT IN ATOMS

Section Review

Objectives

- Describe how to write the electron configuration for an atom.
- Explain why the actual electron configurations for some elements differ from those predicted by the Aufbau principle

Vocabulary

- · electron configurations
- Pauli exclusion principle

Aufbau principle

Huand's rule

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

The ways in which electrons are arranged around the nuclei of atoms are called __1 __. The __2 __ describes the sequence in which orbitals are filled. The various orbitals within a sublevel of a principle energy level are always of __3 __ energy. The __4 __ principle states that a maximum of only __5 __ electrons can occupy each orbital. To occupy the same orbital, two electrons must have __6 __ spins. Hund's rule states that the electrons pair up only after each orbital in a sublevel is occupied by __7 __ When using the shorthand method for showing the electron configuration of an atom, __8 __ are used to indicate the number of __9 __ occupying each sublevel.

Correct electron configurations can be obtained by using the Aufbau diagram for the elements up to and including vanadium. <u>10</u> and copper are exceptions to the Aufbau principle.

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Section 5 2 Electron Arrangement In Atoms Answer Key

Donald A. McQuarrie, Stanley Gill

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