

## Section 5.3 Electron Configuration

In your textbook, read about ground-state electron configurations.

Use each of the terms below just once to complete the passage.

Aufbau principle	electron configuration	ground-state electron configuration	Hund's rule
lowest	Pauli exclusion principle	spins	stable

The arrangement of electrons in an atom is called the atom's

(1) \_\_\_\_\_. Electrons in an atom tend to assume the arrangement that gives the atom the (2) \_\_\_\_\_ possible energy. This arrangement of electrons is the most (3) \_\_\_\_\_ arrangement and is called the atom's (4) \_\_\_\_\_.

Three rules define how electrons can be arranged in an atom's orbitals. The

(5) \_\_\_\_\_ states that each electron occupies the lowest energy orbital available. The (6) \_\_\_\_\_ states that a maximum of two electrons may occupy a single atomic orbital, but only if the electrons have opposite (7) \_\_\_\_\_. (8) \_\_\_\_\_ states that single electrons with the same spin must occupy each equal-energy orbital before additional electrons with opposite spins occupy the same orbitals.

Complete the following table.

Element	Atomic Number	Orbitals					Electron Configuration
		1s	2s	2p <sub>x</sub>	2p <sub>y</sub>	2p <sub>z</sub>	
9. Helium							1s <sup>2</sup>
10.	7						
11. Neon		↑↓	↑↓	↑↓	↑↓	↑↓	

## Section 5 3 Electron Configuration Answer Key

**Bruce C. Gilbert, Damien M.  
Murphy, Victor Chechik**



### Section 5 3 Electron Configuration Answer Key:

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