Stoichiometry Worksheet #1

Aluminum chloride, AlCl₃, is used as a catalyst in various industrial reactions. It is prepared from hydrogen chloride gas and aluminum metal shavings.
 2Al(s) + 6HCl(g) → 2AlCl₃(s) + 3H₂(g)

1A. How many moles of AlCI₂ can be prepared from 3.5 moles of hydrogen chloride gas with an excess of aluminum?

Ans: L2 mol AlCb

1B. How many moles of hydrogen gas would be produced from the use of 8.5 moles of aluminum with an excess of hydrogen chloride?
Ans: 13 mol H₂

2. When dinitrogen pentoxide, N2O2, a white solid, is heated, it decomposes to nitrogen dioxide and oxygen.

$$2N_2O_5(s) \xrightarrow{\Delta} 4NO_2(g) + O_2(g)$$

2A. How many moles of nitrogen dioxide can be formed from the decomposition of 1.25 g of N2Os?

Ans: 0.0231 mol NO:

2B. How many grams of oxygen can be formed from the decomposition of 2.3 g of N₂O₁?

Ans: 0.34 g Oz

Chlorine is prepared from sodium chloride by electrochemical decomposition. Formerly chlorine was produced by heating
hydrochloric acid with pyrolusite [manganese dioxide or manganese(IV) oxide, MnO₂], a common manganese ore. Small amounts
of chlorine may be prepared in the laboratory by the same reaction.

$$4HCl(aq) + MnO_2(s) \rightarrow 2H_2O(l) + MnCl_2(aq) + Cl_2(g)$$

3A. How many grams of HCl react with 5.00 g of MnO2, according to the equation?

Ans: 8.39 g HCI

3B. If a chemist wanted to prepare 100, g of chlorine, how many grams of MnO₂ are needed, assuming there is more than enough hydrochloric acid?
Ans: 123 g MnO₂

3C. How many molecules of water are produced from the reaction of 5.0 g of HCT?

Ans: 4.1 x 1033 molec. H₂O

 Sodium is a soft, reactive metal that instantly reacts with water to give hydrogen gas and a solution of sodium hydroxide, NaOH.

4A. How many grams of sodium metal are needed to give 7.81 g of hydrogen by this reaction? Ans: 178 g Na

4B. How many sodium atoms are needed to react with 1.25 x 10²⁴ molecules of water?

Ans: 1.25 x 1034 atoms Na

Section 9 1 What Is Stoichiometry Answers

John Kenkel, Paul B. Kelter, David S. Hage

Section 9 1 What Is Stoichiometry Answers:

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