

# Stoichiometry: Determining Grams\_\_\_\_\_

Directions: Read and then answer the questions.

Page \_\_\_\_\_

Stoichiometry is the study of the relationships between the amount of reactant used and the amount of product formed by a chemical reaction. Stoichiometry is based on the law of conservation of mass. The law of conservation of mass states that matter is neither created nor destroyed. Thus, in a chemical reaction, the mass of the reactants equals the mass of the products. You can use stoichiometry to answer questions about the amounts of reactants used or products formed by reactions. For example, look at the balanced chemical equation for the formation of table salt (NaCl).



2 sodium atoms react 1 chlorine molecule to form 2 formula units of table salt.

When you look at a balanced equation, there are two ways to interpret the coefficients.

1. The coefficients tell you how many individual particles are interacting in the chemical reaction.
2. The coefficients tell you that two moles of sodium react with one mole of chlorine to form two moles of salt.

The coefficient represents both the number of particles and the number of moles reacting in a chemical reaction.



Coefficients do not represent the mass of the reactants or the mass of products. Knowing the number of moles for each substance allows you to calculate the number of grams for each substance. The mass for each substance is calculated by using molar mass. Remember, 1 mole of any substance is equal to its atomic mass in grams. You can calculate the number of grams for each substance by multiplying the number of moles of that substance by its molar mass.

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# Section Quiz Introduction To Stoichiometry

**Marcel A. Müller**



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