



Interpreting Solubility Curve



1. Define solubility. _____
2. Using the solubility curve below, answer the given questions.



- I. What is the solvent in this solubility curve? _____
- II. Saturation points are given per how much of the solvent? _____
- III. What is the saturation point of NaNO_3 at 10°C , 40°C , & 80°C ?
 10°C = _____ 40°C = _____ 80°C = _____
- IV. How much NH_3 can you dissolve in water at 10°C , 30°C , & 90°C ?
 10°C = _____ 30°C = _____ 90°C = _____
- V. Which salt is the least soluble in water? _____
- VI. How many grams of potassium chloride can be dissolved in 200 g of water at 80°C ? _____
- VII. At 40°C , how much potassium nitrate will be dissolved in 300 g of water? _____
- VIII. Which salt shows the least change in solubility from 0° to 100°C ? _____
- IX. Which substance shows a decrease in solubility from 0° to 100°C ? _____
- X. At 30°C , 90 g of sodium nitrate is dissolved in 100 g water. Is this solution saturated, unsaturated, or supersaturated? _____

Solubility Curves Pogil Answer Key

**Juliette Lantz, Renée Cole, The POGIL
Project**



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Analytical Chemistry Juliette Lantz, Renée Cole, The POGIL Project, 2014-12-31 An essential guide to inquiry approach instrumental analysis Analytical Chemistry offers an essential guide to inquiry approach instrumental analysis collection The book focuses on more in depth coverage and information about an inquiry approach This authoritative guide reviews the basic principles and techniques Topics covered include method of standard the microscopic view of electrochemistry calculating cell potentials the BerriLambert atomic and molecular absorption processes vibrational modes mass spectra interpretation and much more

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