

Name : _____ Date : _____

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Solubility Rules Practice Worksheet

1. Classify each substance as being soluble or insoluble in water.

I. $\text{Mg}(\text{PO}_4)_2$ - _____

VI. KBr - _____

II. KOH - _____

VII. $\text{Pb}(\text{CO}_3)_2$ - _____

III. NiCl_2 - _____

VIII. PbI_2 - _____

IV. NH_4OH - _____

IX. BaSO_4 - _____

V. Hg_2SO_4 - _____

X. NiCl_2 - _____

2. Show the ions that formed the following compounds:

I. $\text{Zn}_3(\text{PO}_4)_2$

II. Al_2S_3

III. Iron (III) sulfide

IV. Ammonium cyanide

3. Form 4 water-soluble compounds by combining ions from the ions below:

Cl^-

CO_3^{2-}

PO_4^{3-}

Li^+

Sr^{2+}

4. Identify the precipitate in the following reaction. Circle the correct answer.



5. Name or give the chemical formula for each of the following compounds and state whether they are soluble (will dissolve) or insoluble (will not dissolve) in solution.

Chemical Formula	Name	Solubility
$\text{Zn}_3(\text{PO}_4)_2$		
	Silver bromide	
KNO_3		
	Aluminum sulfide	
	Silver acetate	

Solubility Rules Practice Answers

Donald A. McQuarrie, Stanley Gill



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Chemistry Theodore Lawrence Brown, 1988

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Table of Contents Solubility Rules Practice Answers

1. Understanding the eBook Solubility Rules Practice Answers
 - The Rise of Digital Reading Solubility Rules Practice Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Solubility Rules Practice Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solubility Rules Practice Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Solubility Rules Practice Answers
 - Personalized Recommendations
 - Solubility Rules Practice Answers User Reviews and Ratings
 - Solubility Rules Practice Answers and Bestseller Lists

5. Accessing Solubility Rules Practice Answers Free and Paid eBooks
 - Solubility Rules Practice Answers Public Domain eBooks
 - Solubility Rules Practice Answers eBook Subscription Services
 - Solubility Rules Practice Answers Budget-Friendly Options
6. Navigating Solubility Rules Practice Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Solubility Rules Practice Answers Compatibility with Devices
 - Solubility Rules Practice Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Solubility Rules Practice Answers
 - Highlighting and Note-Taking Solubility Rules Practice Answers
 - Interactive Elements Solubility Rules Practice Answers
8. Staying Engaged with Solubility Rules Practice Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solubility Rules Practice Answers
9. Balancing eBooks and Physical Books Solubility Rules Practice Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solubility Rules Practice Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Solubility Rules Practice Answers
 - Setting Reading Goals Solubility Rules Practice Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Solubility Rules Practice Answers
 - Fact-Checking eBook Content of Solubility Rules Practice Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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