

Question: What causes bonds to be polar or nonpolar?

1. **Observe:** Select the **Show polar molecule inset** checkbox. The animation shows the probable location of electrons (orange dots) in a polar molecule.
 - A. What do you notice about the distribution of the electrons? I notice they are relatively close to each other.
 - B. How does this electron distribution affect the charges of the bonded atoms? The more electrons, something has, its more negative.
2. **Observe:** Turn on the **Show nonpolar molecule inset**.
 - A. How are the electrons in this molecule distributed? They are evenly distributed.
 - B. Why do the bonded atoms remain neutral? Because neither atom attracts more electrons.
3. **Experiment:** Turn off **Show polar molecule inset** and **Show nonpolar molecule inset**. Check that **Show valence electrons** is turned on. Drag the Na and Cl atoms into the simulation area. Click **Play**. Note the colors. Red indicates a negative charge, while blue indicates a positive charge.
 - A. Does a polar or nonpolar bond form? yes.
 - B. Try several other metal/nonmetal combinations. What do you notice about the bonds that form? I notice that all the bonds between metals and nonmetals form polar bonds.
Ionic bonds are polar bonds that form between metal and nonmetal atoms. In this bond, valence electrons are transferred from a metal to a nonmetal. Drag each of these metal/nonmetal combinations into the ionic bond bin on the upper right.

(Activity A continued on next page)

Activity A (continued from previous page)

4. **Experiment:** Now try forming bonds between different combinations of nonmetals.
 - A. What do you notice? I noticed that the atoms are attracted to the atoms with the greatest electronegativity.
 - B. Are all of these bonds polar? Explain. No, some bonds have difference of EN, making it nonpolar.
- C. Which nonmetal appears to attract electrons the most? Oxygen.
5. **Experiment:** Turn on **Show electronegativity**. **Electronegativity (EN)** describes how strongly an atom attracts a pair of shared electrons. The higher the EN value, the greater the tendency of an atom to hold onto electrons in a bond and become negatively charged. Electronegativity difference is found by subtracting the EN value of one atom from another.
 Choose two nonmetals with a small (or no) EN difference between them. Click **Play**.
 What happens to the shared electrons in this bond? It remains neutral and they aren't really attached to a specific atom.
 When the shared electrons experience the same attraction from each atom, the result is a **nonpolar covalent bond**. Drag the bonded atoms to the **Nonpolar covalent bond** bin.
6. **Experiment:** Choose two nonmetals with a large electronegativity difference. Click **Play**.
 What happens to the valence electrons in this bond? The electrons go to the more negative atom.
A bond in which the electrons are shared unequally is a polar covalent bond. Notice the small $\delta+$ and $\delta-$ symbols, which indicate **partial charges**. Drag the resulting combination to the **Polar covalent bond** bin.
7. **Classify:** Use the Gizmo to categorize the remaining element combinations as forming either ionic, polar covalent, or nonpolar covalent bonds. List your results below.

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Greg Robinson, Mark Rose



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EPA-600/3 , 1991 **Bulletin of the Chemical Society of Japan** Nihon Kagakkai, 1989

Communicating Effectively For Dummies Marty Brounstein, 2011-03-16 A friendly guide that teaches you effective methods of communication to avoid common conflicts and make your voice heard in the office Communicating Effectively For Dummies shows you how to get your point across at work and interact productively with bosses and coworkers Applying your knowledge and skill to your job is the easy part working well with others is often the hard part This helpful guide lets you maximize your personal interactions even when resolving conflicts dealing with customers or giving difficult presentations Whether you re the CEO of a major corporation a small business owner or a team manager effective and clear communication is imperative to your success From keeping your listener engaged to learning to become a better listener Communicating Effectively For Dummies offers all the strategies tips and advice you need to Learn how to become an active listener Accentuate the positive in negative situations Find win win solutions for conflicts Stay on track when writing e mails and letters Handle presentations interviews and other challenges Speak forcefully and assertively without alienating others This friendly and comprehensive guide gives you the keys to a thriving career with expert advice on effective verbal and nonverbal communication From mastering your own facial expressions and reading them in others to being a happy boss this book covers all the angles Becoming aware of your own assumptions Dealing with passive aggressive

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