

Student Exploration: Bohr Model of Hydrogen

Vocabulary: absorption spectrum, Bohr model, electron volt, emission spectrum, energy level, ionization energy, laser, orbital, photon

Prior Knowledge Questions (Answer these questions BEFORE using the Gizmo.)

1. What happens to an electron when it absorbs a tiny packet of light, called a photon?

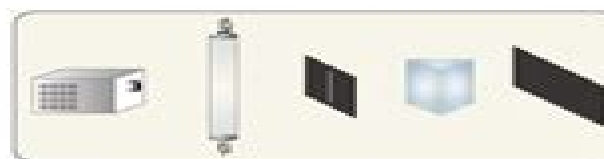
It will move up to a higher energy level.

2. What happens to an electron when it emits a photon?

It will move down to a lower energy level.

Gizmo Warm-up

When white light is passed through hydrogen gas and separated by a prism, some wavelengths of light are absorbed by the hydrogen atoms in the gas. This causes black bars to appear in the **absorption spectrum** of hydrogen. An **emission spectrum** is created when hydrogen gas emits light.



The **Bohr model** of the atom was inspired by the spectrum produced by hydrogen gas. The link between light and atomic structure is illustrated by the *Bohr Model of Hydrogen* Gizmo. The Gizmo shows a **laser** pointed at a container of hydrogen gas. After passing through the gas, light from the laser goes through a prism and is detected on a screen.

1. With the **Laser energy** set to 7.0 eV, click **Play** (▶). Observe the red electron on the **Orbitals** pane. **Do 7.0-eV photons have any effect on the electron?** Photons do not have any effect on the electron as it stays at energy level 1.
2. Click **Pause** (⏸). Set the **Laser energy** to 12.1 eV and click **Play**. **What happens to the electron now?** The electron has moved up to energy level 3.
3. Under **Go to energy level**, select **1**. Notice that a photon is emitted by the electron. **What is the energy of the emitted photon?** 12.1 eV

Student Exploration Bohr Model Of Hydrogen Answer Key

Yicheng Fang



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Table of Contents Student Exploration Bohr Model Of Hydrogen Answer Key

1. Understanding the eBook Student Exploration Bohr Model Of Hydrogen Answer Key
 - The Rise of Digital Reading Student Exploration Bohr Model Of Hydrogen Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Student Exploration Bohr Model Of Hydrogen Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Student Exploration Bohr Model Of Hydrogen Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Student Exploration Bohr Model Of Hydrogen Answer Key

- Personalized Recommendations
 - Student Exploration Bohr Model Of Hydrogen Answer Key User Reviews and Ratings
 - Student Exploration Bohr Model Of Hydrogen Answer Key and Bestseller Lists
5. Accessing Student Exploration Bohr Model Of Hydrogen Answer Key Free and Paid eBooks
 - Student Exploration Bohr Model Of Hydrogen Answer Key Public Domain eBooks
 - Student Exploration Bohr Model Of Hydrogen Answer Key eBook Subscription Services
 - Student Exploration Bohr Model Of Hydrogen Answer Key Budget-Friendly Options
 6. Navigating Student Exploration Bohr Model Of Hydrogen Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Student Exploration Bohr Model Of Hydrogen Answer Key Compatibility with Devices
 - Student Exploration Bohr Model Of Hydrogen Answer Key Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Student Exploration Bohr Model Of Hydrogen Answer Key
 - Highlighting and Note-Taking Student Exploration Bohr Model Of Hydrogen Answer Key
 - Interactive Elements Student Exploration Bohr Model Of Hydrogen Answer Key
 8. Staying Engaged with Student Exploration Bohr Model Of Hydrogen Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Student Exploration Bohr Model Of Hydrogen Answer Key
 9. Balancing eBooks and Physical Books Student Exploration Bohr Model Of Hydrogen Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Student Exploration Bohr Model Of Hydrogen Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Student Exploration Bohr Model Of Hydrogen Answer Key
 - Setting Reading Goals Student Exploration Bohr Model Of Hydrogen Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Student Exploration Bohr Model Of Hydrogen Answer Key

- Fact-Checking eBook Content of Student Exploration Bohr Model Of Hydrogen Answer Key
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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