




Answers to examination-style questions

Answers	Marks	Examiner's tips
1 a) i) 2	1	
ii) 1	1	
iii) 0	1	It is easier to calculate the order with respect to B first in this example.
b) $k = \frac{\text{rate}}{[D]^2[E]}$ or $\frac{8.36 \times 10^{-4}}{(0.84)^2(1.16)}$ $= 1.02 \times 10^{-3} \text{ mol}^{-2} \text{ dm}^6 \text{ s}^{-1}$	3	Make sure you can rearrange formulae. The units are worth a mark so do not leave them out. Answers from 1.02×10^{-3} to 1.05×10^{-3} are acceptable to take account of rounding up in the numbers during the calculation.
2 a) i) Expt 2 2.68×10^{-4} Expt 3 10.72×10^{-4} Expt 4 2.08×10^{-3}	3	Fewer than 2 dp here will be penalised. This gives 3 significant figs in answer. Accept 10.7×10^{-4}
ii) $k = \frac{\text{rate}}{[X]^2}$ or $\frac{2.68 \times 10^{-4}}{(1.20 \times 10^{-3})^2}$ $= 186 \text{ mol}^{-1} \text{ dm}^3 \text{ s}^{-1}$	3	Units are worth a mark so don't leave them out. In this case candidates often do part ii) to calculate k before part i).
b) increases exponentially, i.e. 	1	This is quite hard so a straight line like this was allowed:  but not a curve like this: 

Nelson Thornes Chemistry A2 Answers Chapter 13

E. N. Ramsden



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Table of Contents Nelson Thornes Chemistry A2 Answers Chapter 13

1. Understanding the eBook Nelson Thornes Chemistry A2 Answers Chapter 13
 - The Rise of Digital Reading Nelson Thornes Chemistry A2 Answers Chapter 13
 - Advantages of eBooks Over Traditional Books
2. Identifying Nelson Thornes Chemistry A2 Answers Chapter 13
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Nelson Thornes Chemistry A2 Answers Chapter 13
 - User-Friendly Interface
4. Exploring eBook Recommendations from Nelson Thornes Chemistry A2 Answers Chapter 13
 - Personalized Recommendations
 - Nelson Thornes Chemistry A2 Answers Chapter 13 User Reviews and Ratings
 - Nelson Thornes Chemistry A2 Answers Chapter 13 and Bestseller Lists
5. Accessing Nelson Thornes Chemistry A2 Answers Chapter 13 Free and Paid eBooks
 - Nelson Thornes Chemistry A2 Answers Chapter 13 Public Domain eBooks
 - Nelson Thornes Chemistry A2 Answers Chapter 13 eBook Subscription Services
 - Nelson Thornes Chemistry A2 Answers Chapter 13 Budget-Friendly Options
6. Navigating Nelson Thornes Chemistry A2 Answers Chapter 13 eBook Formats

- ePub, PDF, MOBI, and More
 - Nelson Thornes Chemistry A2 Answers Chapter 13 Compatibility with Devices
 - Nelson Thornes Chemistry A2 Answers Chapter 13 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Nelson Thornes Chemistry A2 Answers Chapter 13
 - Highlighting and Note-Taking Nelson Thornes Chemistry A2 Answers Chapter 13
 - Interactive Elements Nelson Thornes Chemistry A2 Answers Chapter 13
 8. Staying Engaged with Nelson Thornes Chemistry A2 Answers Chapter 13
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Nelson Thornes Chemistry A2 Answers Chapter 13
 9. Balancing eBooks and Physical Books Nelson Thornes Chemistry A2 Answers Chapter 13
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Nelson Thornes Chemistry A2 Answers Chapter 13
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Nelson Thornes Chemistry A2 Answers Chapter 13
 - Setting Reading Goals Nelson Thornes Chemistry A2 Answers Chapter 13
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Nelson Thornes Chemistry A2 Answers Chapter 13
 - Fact-Checking eBook Content of Nelson Thornes Chemistry A2 Answers Chapter 13
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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