



Name: \_\_\_\_\_ (Surname First) \_\_\_\_\_ Date: \_\_\_\_\_

## Student Exploration: Colligative Properties

**Background:** Colligative properties are properties depending on concentration. Colligative properties include boiling point, freezing point, osmotic pressure, vapor pressure, boiling point elevation, freezing point depression, and solvent potential.

**Prior Knowledge:** Questions (for this and all Student Explorations):

- What is the difference between a homogeneous and a heterogeneous mixture? Give an example of each. (Do you know?)
- Boiling point is the temperature at which a liquid changes to a gas. What does it depend on?

- Antifreeze is added to water to prevent freezing. How does it do this? Antifreeze's boiling point is higher than water's. How does this affect boiling point? (Do you know? Antifreeze's boiling point is higher than water's.)

Antifreeze makes the boiling point of water go up. Antifreeze has a higher boiling point than water, so it doesn't evaporate as easily. This means that the liquid stays at a higher temperature for longer. This is why it's used in car engines to keep them cool.

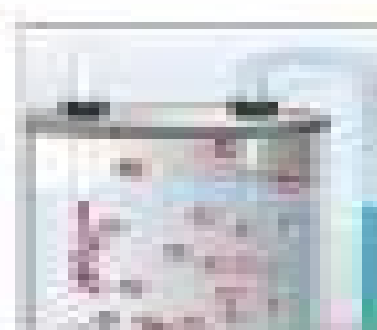
**What You Will Do:**

Working with an adult supervisor, you will make solutions for which you determine boiling points. At least 10 solutions are made. These solutions and others, known as <https://www.gimnos.com>, are compared at this Colligative Properties lab.

There are three parts to this lab: making the solutions, boiling them, and comparing the results.

- What is the difference between a homogeneous and a heterogeneous mixture? Give an example of each. (Do you know?)

Boiling happens at a certain temperature and pressure. This is why it's used in cooking and in many other processes.



# Student Exploration Colligative Properties Gizmo

## Answers

**S Nieto**



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