

Electrons & the Periodic Table – SOL review packet 2

(ALL NUMERICAL ANSWERS MUST INCLUDE THE PROPER UNITS)

NAME THE TERM DESCRIBED BY THE FOLLOWING DEFINITIONS:

1. The progression of positions of an element in the lowest possible energy levels.
2. When an electron jumps up to a higher energy level, the atom is in its _____.
3. The potential energy applied (Bohr's) positive energy theory to electrons.
4. The theory that it is impossible to know both the position and speed of an electron simultaneously.
5. The theory that no electron in an atom can share the same 4 quantum numbers.
6. The theory that electrons in the lowest energy orbitals first.
7. The theory that within a sublevel, electrons prefer to occupy their own orbitals.
8. A term describing the sublevel electrons in an atom.
9. A positively charged ion is called with _____, a negatively charged ion is called with _____.
10. The most stable type of electron configuration.
11. The lower left hand section of the periodic table that includes most of the elements.
12. The upper right hand section of the periodic table.
13. The section of the periodic table that includes only the elements in the p-block.
14. The section of the periodic table that includes both the s-block and the p-block.
15. A three-dimensional region in space where an electron is likely to be found.
16. The process by which the core electrons partially block the attraction between the nucleus and the valence electrons.

WRITE CORRECT SYMBOLS FROM THE FOLLOWING ELEMENTS:

17. Mg
18. Si
19. Fe

WRITE CORRECTED ELECTRON CONFIGURATIONS FROM THE FOLLOWING ELEMENTS:

20. Fe
21. Al
22. Cr

WRITE CORRECTED ELECTRON CONFIGURATIONS FROM THE FOLLOWING ELEMENTS:

23. Sn

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