

1. Some of the energy levels of the hydrogen atom are shown below. (consider all of the level jumps)

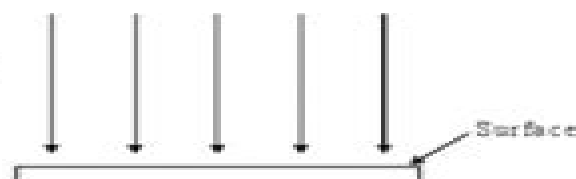


Electrons are excited to the -0.85 eV level. How many different photon frequencies will be observed in the emission spectrum of hydrogen?

- a. 3 b. 4 c. 5 **d. 6**

2. Light of wavelength $6.0 \times 10^{-7} \text{ m}$ is incident normally on a plane surface as shown below.

Incident light, wavelength $6.0 \times 10^{-7} \text{ m}$



The light photons are absorbed by the surface.

Determine the energy for one photon of light. ($h = 6.63 \times 10^{-34} \text{ Js}$ and $c = 3.00 \times 10^8 \text{ m/s}$)

- a. $2.2 \times 10^{-28} \text{ J}$ **b. $3.3 \times 10^{-19} \text{ J}$** c. $4.5 \times 10^{-48} \text{ J}$ d. $1.6 \times 10^{-28} \text{ J}$

$$E = \frac{hc}{\lambda} = \frac{(6.63 \times 10^{-34})(3 \times 10^8)}{6.0 \times 10^{-7}} = 3.3 \times 10^{-19} \text{ J}$$

Unit Vii Quiz 1 Answers Physics

Emilie Sanchez

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name bar.

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