

## Limiting Reagents and Percentage Yield Worksheet

### 1. Consider the reaction



- a) 80.0 grams of iodine(V) oxide,  $\text{I}_2\text{O}_5$ , reacts with 28.0 grams of carbon monoxide,  $\text{CO}$ . Determine the mass of iodine  $\text{I}_2$ , which could be produced?

80 g $\text{I}_2\text{O}_5$	1 mol $\text{I}_2\text{O}_5$	1 mol $\text{I}_2$		XS
1	333.8 g $\text{I}_2\text{O}_5$	1 mol $\text{I}_2\text{O}_5$		
28 g $\text{CO}$	1 mol $\text{CO}$	1 mol $\text{I}_2$	253.8 g $\text{I}_2$	
1	28 g $\text{CO}$	5 mol $\text{CO}$	1 mol $\text{I}_2$	= 50.8 g $\text{I}_2$

- b) If, in the above situation, only 0.160 moles, of iodine,  $\text{I}_2$  was produced.  
i) what mass of iodine was produced?

0.160 mol $\text{I}_2$	253.8 g $\text{I}_2$	= 40.6 g $\text{I}_2$
	1 mol $\text{I}_2$	

- ii) what percentage yield of iodine was produced.

40.6 g $\text{I}_2$	$\times 100 =$	79.9 %
50.8 g $\text{I}_2$		

### 2. Zinc and sulphur react to form zinc sulphide according to the equation.



If 25.0 g of zinc and 30.0 g of sulphur are mixed,

- a) Which chemical is the limiting reactant?

25.0 g Zn	1 mol Zn	1 mol ZnS	= 382 mol ZnS	LR
1	65.41 g Zn	1 mol Zn		
30.0 g S	1 mol S	1 mol ZnS	= 934 mol ZnS	XS
1	32.1 g S	1 mol S		

- b) How many grams of  $\text{ZnS}$  will be formed?

382 mol ZnS	97.51 g ZnS	= 37.2 g ZnS
	1 mol ZnS	

- c) How many grams of the excess reactant will remain after the reaction is over?

$$.934 \text{ mol ZnS} - .382 \text{ mol ZnS} = .552 \text{ mol ZnS}$$

.552 mol ZnS	1 mol S	32.1 g S	= 17.7 g S
	1 mol ZnS	1 mol S	

### 3.

# Unit 4 Limiting Reactant And Percent Yield Ws 7

**Michal Rosen-Zvi**



#### **Unit 4 Limiting Reactant And Percent Yield Ws 7:**

**Theoretical Chemical Engineering Abstracts** ,1980      **Limiting Reactant and Percent** Lifeliqe,2019 This lesson plan covers analyzing chemical reactions in order to determine limiting reactants and excess reactants including calculating the amount of excess reactant calculate the theoretical yield of a reaction and calculating the percent yield of a reaction

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