

## Answers to Worksheet 6.1 Writing and Balancing Equations

1. Write the chemical equations and balance each of the following word equations.

- a) Aluminum metal reacts with iron (II) oxide powder to produce aluminum oxide solid and iron metal.



- b) Aluminum sulphate solution and calcium hydroxide solution produce a precipitate of aluminum hydroxide and solid calcium sulphate.



- c) Ammonia gas ( $\text{NH}_3$ ) plus oxygen gas yields nitrogen (I) oxide gas plus water vapour.



- d) Calcium hydroxide solution and carbon dioxide gas yields solid calcium carbonate and liquid water.



- e) Aqueous iron (III) chloride and sodium carbonate solution yields aqueous sodium chloride and a precipitate of iron (III) carbonate.



- f) Solid iron (III) oxide and carbon monoxide gas yields iron metal and carbon dioxide gas.



- g) Magnesium carbonate solution plus aqueous hydrochloric acid (HCl) yields magnesium chloride solution plus liquid water and carbon dioxide gas.



- h) Silicon dioxide solid plus aqueous hydrofluoric acid (HF) yields solid silicon tetrafluoride plus liquid water.



- i) Aqueous sodium hydroxide and carbon dioxide gas yields sodium carbonate solution and liquid water.



# Unit 6 Chemical Reactions Answers

**Thomas Bond, Chris Hughes**



## Unit 6 Chemical Reactions Answers:

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