

Name _____ Date: _____ Class period _____



Unit: Periodic Trends "Atomic Size Trend" - Wksh #2



Directions: Please answer each fill in the blank with the best answer. ☹

1. As you move from top to bottom down a group on the periodic table the size of an atom will _____.
2. This happens because as you move down the group each element has one more occupied _____ than the last one.
3. As you move from left to right across a period on the periodic table the size of an atom will _____.
4. This happens because as you move across a period each element has one more _____ in its nucleus than the last one.
5. Within a period the elements will have the same number of occupied _____.
6. The _____ effect is caused by inner electrons interfering with the attraction of outer electrons for the protons in the nucleus.
7. Electrons which are further from the nucleus are _____ attracted to the nucleus because of this effect.
8. The decrease in atomic size as you move from left to right across the periodic table becomes less pronounced as you move _____ the periodic table.
9. Explain why fluorine has a smaller atomic radius than oxygen.
10. Explain why fluorine has a smaller atomic radius than chlorine.
11. Arrange these elements in order of decreasing atomic size: sulfur, chlorine, aluminum, and sodium. Does your arrangement demonstrate a periodic or a group trend?

Circle the element in each of the following pairs with the greater atomic size:

- | | |
|-------------------------|-----------------------|
| a) sodium or lithium | c) carbon or boron |
| b) silicon or magnesium | d) selenium or oxygen |

Unit 8 Quiz Periodic Trends Answer Key

Disha Experts

A red circular graphic with a gradient, appearing as a partial circle or a thick curved line, located to the right of the 'Disha Experts' text.

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